

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
8 March 2001 (08.03.2001)

PCT

(10) International Publication Number
WO 01/17052 A2

(51) International Patent Classification⁷: **H01M 10/40**,
H01G 9/02, H01B 1/12

(21) International Application Number: PCT/US00/22917

(22) International Filing Date: 21 August 2000 (21.08.2000)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
09/388,733 2 September 1999 (02.09.1999) US

(71) Applicant: **LITHIUM POWER TECHNOLOGIES, INC.** [US/US]; 20955 Morris Avenue, Manvel, TX 77578 (US).

(72) Inventor: **MUNSHI, M., Zafar, A.**; P.O. Box 978, Manvel, TX 77578 (US).

(74) Agents: **MINTZ, Carol, E. et al.**; Conley, Rose & Tayon, P.C., P.O. Box 3267, Houston, TX 77253-3267 (US).

(81) Designated States (*national*): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CR, CU, CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, UZ, VN, YU, ZA, ZW.

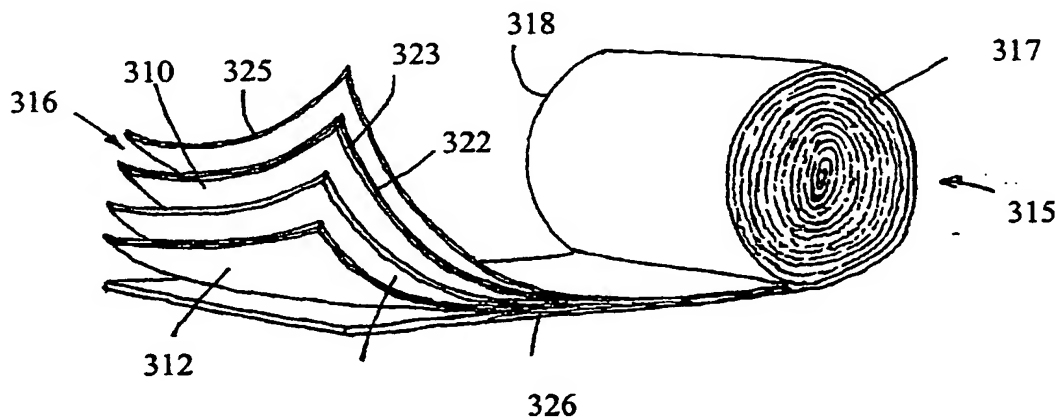
(84) Designated States (*regional*): ARIPO patent (GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG).

Published:

— Without international search report and to be republished upon receipt of that report.

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: **ALL-SOLID-STATE ELECTROCHEMICAL DEVICE AND METHOD OF MANUFACTURING**



(57) Abstract: All-solid-state electrochemical cells and batteries employing very thin film, highly conductive polymeric electrolyte and very thin electrode structures are disclosed, along with economical and high-speed methods of manufacturing. A preferred embodiment is a rechargeable lithium polymer electrolyte battery. New polymeric electrolytes employed in the devices are strong yet flexible, dry and non-tacky. The new, thinner electrode structures have strength and flexibility characteristics very much like thin film capacitor dielectric material that can be tightly wound in the making of a capacitor. A wide range of polymers, or polymer blends, characterized by high ionic conductivity at room temperature, and below, are used as the polymer base material for making the solid polymer electrolytes. The preferred polymeric electrolyte is a cationic conductor. In addition to the polymer base material, the polymer electrolyte compositions exhibit a conductivity greater than 1×10^{-4} S/cm at 25 °C or below and contain an electrically conductive polymer, a metal salt, a finely divided ionic conductor, and a finely divided inorganic filler material. Certain rechargeable batteries of the invention provide high specific energy (250 to 350 Wh/kg) (gravimetric) and energy density (450 to 550 Wh/l) (volumetric), high cycle life (1000 cycles), low self-discharge and improved safety.

WO 01/17052 A2

ALL-SOLID-STATE ELECTROCHEMICAL DEVICE AND METHOD OF MANUFACTURING

CROSS-REFERENCE TO RELATED APPLICATIONS

5 This application is related to U.S. Pat. App. No. 09/388,741, filed September 2, 1999, entitled "Solid Polymer Electrolytes," the disclosure of which is incorporated herein by reference.

BACKGROUND OF THE INVENTION

Field of the Invention

10 The present invention generally relates to the manufacture of solid state electrochemical cells, particularly high energy density cells having very thin electrode and electrolyte structures for building ultra-thin laminar batteries. The invention also relates to lithium polymer electrolyte batteries containing such electrochemical cells and to their methods of making.

Description of the Related Art

15 Lithium rechargeable batteries are the focus of intense investigation around the world because of the rapid proliferation of portable electronic devices in the international marketplace. The state-of-the-art lithium battery is a lithium ion battery which uses a carbon electrode as the negative electrode or anode and a lithiated metal oxide, such as lithiated cobalt oxide, lithiated nickel oxide, lithiated manganese oxide, or mixtures of these materials as the positive electrode
20 or cathode, a microporous polypropylene or polyethylene separator that separates the two electrodes and prevents them from shorting electrically, and liquid organic solvents containing a lithium salt as the electrolyte. The electrolyte is usually absorbed into the separator material and provides high ionic conductivity (10^{-3} to 10^{-2} S/cm) and migration of ions between the electrodes of the cell. These batteries are commercially available and are commonly used in portable
25 computers, cellular telephones and camcorders among other applications. The specific energy and energy density of the lithium ion battery is usually about 125 Wh/kg and 260 Wh/l. Usually, the packaged battery (in a hard plastic case) has a much lower energy density than the individual cell (20% lower). The cycle life (number of times the battery can be recharged) of this battery is about 500 to 800 cycles, the self-discharge (i.e. loss of capacity on standing) per
30 month is about 10%, and cost is about \$1 per Watt-hour of energy. These batteries can be manufactured at a high rate of speed. Even though this battery technology is being commercialized very heavily, there have been numerous safety questions. Cells that have been abused under crush test or high temperature test have been known to explode and catch fire.

An offshoot of the above system is the lithium ion polymer electrolyte battery. The electrode chemistry is the same, but the liquid electrolyte (up to 70% by weight of the electrolyte) in this case is absorbed in a polymer membrane instead of the microporous polypropylene separator. It is based on the Bellcore's U.S. Patent No. 5,296,318 utilizing a polyvinylidene fluoride (PVDF) polymer as the medium that absorbs the electrolyte solvent. Ironically, PVDF is non-conducting and so its sole function is to hold the liquid organic solvent(s) in its structure in a manner similar to a sponge holding water. Because the technology uses an electrolyte solvent absorbed in a polymer, it is not easy to manufacture cells in high speed. Automation of this technology may be very difficult. It is believed the energy densities (gravimetric and volumetric) for this type of battery are lower than the existing lithium ion batteries, cycle life is not too impressive, and cell cost is several dollars per Watt-hour. The physical forms the lithium ion polymer cell could take was a heavily touted feature, but today only flat prismatic cells are typically manufactured.

Another rechargeable lithium battery system uses a lithium metal as the negative electrode instead of carbon, and so the energy density of this system can be increased tremendously because of the very high specific capacity of metallic lithium compared to carbon. Gravimetric specific energy density as high as 200 to 250 Wh/kg have been reported in the literature for rechargeable lithium metal batteries. Lithium metal anode batteries are not new. Primary, non-rechargeable batteries using lithium metal, non-aqueous organic electrolytes and a positive electrode have been used in many applications for the past 25 years despite the fact that lithium metal is thermodynamically unstable in liquid organic solvents and reacts with the solvent irreversibly. Secondary batteries using lithium metal as the negative electrode, intercalation or insertion compounds as the positive electrode, and non-aqueous organic electrolytes were the focus of intense investigation during the 1970's and 1980's. However, the problem with using lithium in a rechargeable system is that because of the instability of lithium in these solvents, a large excess of lithium is required to off-set the chemical reaction of lithium with the solvent, usually as much as 3 to 5 fold. In addition, the cycle life of lithium metal batteries in organic solvent electrolytes is less than 200 cycles. Lithium plating and stripping during the charge and discharge cycles creates a porous deposit of very high surface area and increased activity of the lithium metal with respect to the electrolyte. The reaction is highly exothermic and the cell can vent with flame if heated or short-circuited. Much effort has been expended to improve the cycling efficiency of the lithium anode through changes to the electrolyte or investigating alloys of lower lithium reactivities. Safety features such as fusible separators which cease the electrochemical reactions when the battery temperature approaches a

critical value and overcharge protection redox couples have also been incorporated to improve the safety of these cells.

In addition, liquid solvent electrolyte in any of the above cell systems is often corrosive and toxic and presents handling difficulties through spillage or leakage from the cell. It can also outgas during overcharge or overdischarge or at elevated temperatures, leading to safety problems. Most of the problems have been associated with the electrode/electrolyte interface.

In order to overcome the disadvantages inherent in liquid electrolytes and to obtain superior long-term storage stability there is interest in solid polymeric electrolytes in which ion mobility is achieved through coordination by sites on the polymer chain of electrolyte ions, thus promoting electrolyte dissolution and salt dissociation. An all-solid-state battery using an ionically conductive polymer membrane as the electrolyte would have several attractive features. It could be made into virtually any shape and size, be reasonably rugged and leakproof, and would have low self-discharge. It could be made into thin film power cells or thick film energy cells, would have high open-circuit potentials using a lithium anode, and could be produced by automated fabrication techniques. These features represent a unique combination of properties and give rise to the possibility of using such batteries, as either secondary or primary devices across a wide range of applications.

One polymer that has been examined extensively for use in a solid state battery is poly(ethylene oxide) or PEO, which is able to form stable complexes with a number of salts. Because of its low ionic conductivity at ambient temperature of about 10^{-9} to 10^{-8} S/cm, batteries examined using this material had to operate at 100 °C and above. A major problem with PEO based electrolytes at temperatures below 60°C is their high crystallinity and the associated low ion mobility. In recent years a number of radically different approaches have been taken to improve the conductivity of PEO and PEO-based polymers that have also led to the proposal of other polymers. These approaches included, polymer modifications and synthesizing new polymers; forming composite polymers with ceramic materials; using plasticizer salts to increase the ion transport and mobility of the cation; using plasticizing solvents in the polymer again to increase the ionic character of the cation; among other approaches. Several review articles describe these approaches in detail, e.g. "Technology Assessment of Lithium Polymer Electrolyte Secondary Batteries" by M. Z. A. Munshi, Chapter 19 in Handbook of Solid State Batteries and Capacitors, Ed. M. Z. A. Munshi (World Scientific Pub. Singapore) 1995; A. Hooper, M. Gauthier, and A. Belanger, in: "Electrochemical Science and Technology of Polymers - 2, Ed. R. G. Linford (Elsevier Applied Science, London), 1987.

Polymer modification and synthesis of new polymers resulted in some improvement in the ionic conductivity but the mechanical property and integrity were poor. Probably, the best known polymer as a result of this synthesis is poly(bis(methoxyethoxyethoxide))-phosphazene, known as MEEP, which has an ionic conductivity of about 10^{-5} S/cm at room temperature when combined with a lithium salt, but which has mechanical properties similar to glue. On the other hand, materials based on blocked copolymers may provide alternatives. For example, PEO-PPO-PEO crosslinked with trifunctional urethane and a lithium salt has an ionic conductivity of about 10^{-5} S/cm but is too rigid, brittle and difficult to manufacture.

Inorganic conducting and non-conducting fillers have also been used to increase the ionic conductivity and mechanical property of the polymer. Addition of alpha alumina to (PEO)₈.LiClO₄ resulted in a negligible effect on the ionic conductivity but dramatically increased the mechanical property at 100°C, while the addition of other ceramic materials such as ionically conductive beta alumina to PEO-NaI and PEO-LiClO₄ complexes improved the ionic conductivity of PEO based electrolytes to about 10^{-5} S/cm. In another battery technology, inorganic fillers based on high surface area alumina and silica have been used to enhance the ionic conductivity of lithium iodide solid electrolyte from 10^{-7} S/cm to $10^{-5} - 10^{-4}$ S/cm at room temperature (see C.C. Liang, *J. Electrochemical Society*, Vol. 120, page 1289 (1973)). Plasticizer salts based on lithium bis(trifluoromethane sulfonyl) imide, LiN(CF₃SO₂)₂ trademarked as LiTFSI by Hydro-Quebec and distributed by the 3M Company under the product name, HQ-115 when added to PEO yields a conductivity of about 10^{-5} S/cm.

None of the previous approaches toward improving polymer conductivity has resulted in adequate conductivity enhancements of the polymer electrolytes to permit room temperature operation of batteries utilizing the electrolyte. Accordingly, an attempt was made to increase the ionic conductivity of PEO-based polymer electrolyte by incorporating plasticizing solvents or low molecular weight polymers to the polymer electrolyte. The intent was to increase the ionic mobility and concentrations of the charge carriers in the solid polymer electrolyte by enhancing the dissociation of the lithium salt. Generally, it is believed that the lithium ion is also solvated to the solvent molecule and participates in enhancing the ionic mobility. Many electrolyte composites incorporating low molecular weight polymers or liquid organic solvents have been prepared and have demonstrated high conductivity at room temperature approaching conductivity values of the typical non-aqueous liquid electrolytes. For example, Kelly et al. (*J Power Sources*, 14:13 (1985)) demonstrate that adding 20 mole percent of liquid polyethylene glycol dimethyl ether polymer (PEGDME) to solid PEO polymer results in an increase in the

ionic conductivity of the final plasticized polymer from 3×10^{-7} S/cm to 10^{-4} S/cm at 40°C. However, the mechanical property of this material was very poor.

Bauer et al in U.S. Patent No. 4,654,279 (1987) demonstrate that thermal crosslinking of polymers consisting of epoxies and methacrylates and plasticized with a solution of LiClO₄ in a 400 MW PEG resulted in a conductivity of 4×10^{-4} S/cm at 25°C. This patent describes a polymeric electrolyte consisting of a two phase interpenetrating network (IPN) of a mechanically supporting phase of a continuous network of a cross-linked polymer and an ionically conducting phase comprising of a metal salt and a liquid polymer such as liquid PEG.

Many of these low molecular weight polymers have a relatively low dielectric constant when compared to their liquid solvent counterpart, and thus limit the number of charge carriers in the plasticized polymer. In an effort to overcome this hindrance, high dielectric constant liquid organic solvent such as ethylene carbonate (EC) and propylene carbonate (PC) have been incorporated in the host polymer both to increase the number of charge carriers and increase further the room temperature conductivity of the polymer. The use of these organic solvents to plasticize polymers such as poly(vinyl acetal), poly(acrylonitrile), poly(vinyl acetate) and hexafluoropropene-vinylidene fluoride copolymer (Viton™) were made as early as 1975 by Feuille and Perche (*Journal of Applied Electrochemistry*, Vol. 5, page 63 (1975)). However, the mechanical properties of these polymers were so poor that they had to be supported on porous matrices. Later Armand (*Proc. Workshop on Li Non-Aqueous Battery Electrochemistry, The Electrochemical Soc.* Vol. 80-7, page 261 (1980)) by crosslinking Viton™ and plasticizing with a solution of 1M LiClO₄ in PC produced a system with good room temperature conductivity (10^{-4} S/cm) and good mechanical properties. Polyvinylidene fluoride (PVDF) and polyacrylonitrile (PAN) were evaluated in the early 1980s and have also been doped with a variety of liquid polar solvents, yielding room temperature conductivities as high as 10^{-3} S/cm. Subsequently, PVDF has been the subject of a recent patent from Bellcore (U.S. Patent No. 5,296,318).

The use of PC in an ionically conductive matrix containing oxygen donor atoms such as PEO complexed with a lithium salt was first presented by the present inventor in a paper presented at the Fall Meeting of the Electrochemical Soc. held October 18-23, 1987). Although room temperature battery performance data was presented at that time, the propylenecarbonate/lithium salt/polymer electrolyte did not have good mechanical properties. In the late 1980s through early 1990s, a series of U.S. patents including Nos. 4,792,504; 4,747,542; 4,792,504; 4,794,059; 4,808,496; 4,816,357; 4,830,939; 4,861,690; 4,925,751;

4;925,752; 4,935,317; 4,960,655; 4,990,413; 4,997,732; 5,006,431; 5,030,527; 5,057,385;
5,066,554 and European patents EP 0 359 524 and EP 0 411 949 were issued variously to MHB
Inc. and Hope Industries. These patents described predominantly radiation curing methods for
the preparation of interpenetrating polymeric networks containing various types of polyacrylates
and liquid organic solvents. Although electron beam curing was the preferred method for
polymerizing the IPN, thermal and ultraviolet curing methods were also proposed. The idea
behind this was to contain the PC solution in the matrix of the polymeric network that would
therefore yield a high ionic conductivity comparable to that of PC itself. Indeed, this was
demonstrated in typical polymeric networks, yielding conductivities of about 2×10^{-3} S/cm at
room temperature. An advantage with using electron beam curing compared to UV radiation is
that an electron beam can penetrate through metallic components, and hence complete prototype
cells can be made *in-situ*.

While the addition of organic plasticizers may help solve the problem of low ionic
conductivity in polymer electrolytes, they necessarily introduce additional electrolyte
components that may have deleterious effects on other electrolyte properties, such as stability in
contact with metallic lithium. Like the liquid organic electrolytes, plasticized polymer
electrolyte is not thermodynamically stable at the lithium metal potential. In addition, polymer
electrolytes based on such designs cannot be manufactured in very thin film form so as to reduce
their overall resistance and hence cell resistance, since the polymer will not have the sufficient
strength to hold the liquid organic solvents in its matrix. For such a system to be fully functional
it must be based on a thick film concept, which increases its overall cell resistance and reduces
the energy density due to a reduction in the active components in the cell. Another problem with
this type of design is the fact that polymers containing liquids cannot be wound along with the
rest of the electrode components in a winding machine, since the liquid will tend to ooze out of
the polymer as soon as any stress is applied to the polymer.

Advanced commercial battery technologies such as the lithium ion (Sony type),
THINFILMTM lead-acid (Bolder Battery Co.), and nickel-metal hydride batteries do not
demonstrate very high cycle life, i.e., several tens of thousands of cycles. This is due
primarily to considerable phase changes and composite structure segregation that take place
within the electrode, inefficient materials utilization and particle movements, and expansion
and contraction during charge and discharge. However, batteries such as these do represent a
breakthrough in the design of lithium ion batteries and yielded higher cycle life than was
previously available by their use of relatively thin electrode structures. The THINFILMTM
lead-acid battery for high power applications is a good example of thin electrode structures

and good materials utilization, but the electrode is still not thin enough to achieve a cycle life of several thousand. Hence, a major drawback of these batteries is that the liquid electrolytes necessarily limit their cycle life.

Lithium primary battery electrodes are traditionally made by calendering the cathode paste onto a nickel or stainless steel gauze substrate material and compacting between heated rollers. In the case of lithium metal anodes the gauze is used as a substrate material. The substrate material is usually about 2 to 3 mils thick while the anode and cathode are typically about 5 to 10 mils thick. The anode and cathode are sandwiched with a microporous polypropylene separator between them and wound in a jelly-roll manner. Usually, the laminates are very thick and the electrode length is about two feet in a typical AA size cell. Rechargeable lithium metal anode batteries, which were available a few years ago, were also constructed in this manner.

These techniques have changed considerably with the advent of lithium ion battery construction. In lithium ion batteries the carbon anode is pasted in relatively thin film form onto a copper foil electrode and the lithiated metal oxide cathode is pasted onto an aluminum foil. The substrate thickness for both anode and cathode is in a range from about 25 to 35 microns, and the active electrode is about 25 microns thick. Additionally, the length of each electrode in a typical AA size cell is about twice that of lithium anode cells. Present day electrode/electrolyte component thickness in gelled electrolyte lithium ion cells is of the order of 50 to 75 microns each. This remains far too thick for optimum electrode utilization and high rate capability. Metallic current collectors are also used, not only adding weight but unwanted thickness to the battery. The thick electrode concept in commercial cells is designed for maximum capacity, while the thick gelled PVDF electrolyte provides ease of handling. However, the internal resistance of this battery is still relatively higher than its liquid electrolyte counterpart, thus decreasing battery performance. Cells constructed from such a design cannot be used at high discharge and charge rates. Thick inactive substrates used in such cell construction effectively reduce the energy density of the battery. In addition, this design exposes the cells to risk of high polarization during charge and discharge, which could lead to breakdown of the liquid solvent electrolyte and consequently loss of capacity, loss of cycle life and inadequate safety.

On the other hand, an existing lead acid battery designed for an electric car can only provide about 60 to 80 miles range before recharge is not suitable either. Known lithium ion batteries are neither cost-effective nor safe for electric vehicles. Furthermore, consumers are generally unwilling to pay a premium price for an electric vehicle while gasoline vehicles

remain far less expensive to operate. However, if the price of the battery can be reduced, and the mileage range of the electric vehicle can be increased to a value closer to the mileage range of a conventional vehicle, the electric vehicle will become quite feasible for widespread use.

5 Another problem with existing rechargeable batteries, even those used in small electronic devices, is that they contain liquid organic solvents that have been found to be relatively unsafe for consumer use. Also, there have been many reports of safety problems with lithium ion batteries, in particular, containing organic solvents as the electrolyte in portable electronics devices. The lithium ion batteries must have special charging circuitry since the
10 battery cannot tolerate any amount of overcharge. Also, they have no built-in chemical mechanism, such as that in nickel-cadmium batteries, that provides for recombination reactions to take over when cells reach full charge, so that the electrolyte never decomposes. In liquid organic electrolyte cells, there is no recombination mechanism and the cell electrolyte quickly start to decompose with overvoltage, leading to production of unwanted gaseous species.

15 Much of the literature in this field reports that polymer electrolyte batteries can be made in very thin film form and can be flexible. However, batteries containing the polymer electrolytes with liquid organic solvents actually tend to lose performance over time, primarily because of cell orientation related problems. A battery standing upright will tend to have the liquid solvents travel to the bottom of the cell, and during charge and discharge, the current
20 along the cell height will be different because of the difference in the conductivity at the bottom of the electrode and at the top. Such cells do not have very high cycle life and lose capacity as a result of poor charging and discharging.

Glassy electrolytes based on the ternary system, $\text{LiI-Li}_2\text{S-P}_2\text{S}_5$ and $\text{LiI-Li}_2\text{S-B}_2\text{S}_3$ exhibit lithium ion conductivities of about 1×10^{-3} S/cm and 2×10^{-3} S/cm at room temperature,
25 respectively, for bulk electrolytes. Since the electrolyte is very hard and brittle, it is usually used in a vapor deposited form in very thin films so as to achieve a continuous uniform thickness. Lithium anode batteries based on thin film glassy electrolytes with electrolyte thickness of one to several microns have demonstrated current densities of several tens to hundreds of microamperes per square centimeter at room temperature. Unfortunately, the fabrication of
30 these batteries requires extremely sophisticated deposition equipment and production feasibility and performance of large cells have not yet been demonstrated.

All of the prior art techniques that have been employed to improve the ionic conductivity, mechanical strength, safety, and chemical stability, and to reduce cost by simplifying or improving the synthesis of polymer electrolytes have one or more serious

shortcomings. As a result, there is still no room temperature conducting polymer electrolyte available today that is entirely suitable for use with a lithium metal anode in a rechargeable lithium battery, for example. There remains a need for electrochemical cells with polymer electrolytes that have good ionic conductivity at room temperature and below, so that the performance of an electrochemical cell at room temperature or below can be improved. There is also a great need for thinner component batteries that avoid the use of organic solvents in the electrolyte without sacrificing energy density.

SUMMARY OF THE INVENTION

The invention is as described in the claims. New all-solid-state laminar electrochemical cells, anodes, cathodes, electrodes and half-electrodes, bipolar units, batteries and methods according to the present invention overcome many of the difficulties and disadvantages associated with conventional electrochemical devices and their methods of making. Certain preferred electrochemical cells and batteries of the present invention avoid many of the kinetic constraints on the specific power, cycling efficiency and capacity utilization that are typical of electrochemical devices employing conventional polymer electrolytes and electrodes. A kinetic constraint is, for example, the inability to discharge at high rates leading to polarization losses. The lithium solid polymer electrolytes employed in the cells and batteries of certain preferred embodiments of the present invention are suitable for use at room temperature, and even below, due to their high ionic conductivity compared to previous lithium polymer electrolytes. Not only are these new electrolytes highly conductive, they are also strong enough to be manufactured by high speed roll-to-roll techniques. One advantage of the preferred anodes, cathodes, electrodes and half-electrodes, bipolar units, electrochemical cells and batteries of the present invention is that they employ a thermodynamically stable dry polymer electrolyte that can be manufactured using high speed extrusion or deposition techniques, according to certain preferred methods of the invention.

While the secondary behavior of existing lithium batteries has been limited by a number of factors related to the materials and design of the system, the preferred ultra-thin battery embodiments of the present invention overcome those drawbacks by incorporating very thin electrode and electrolyte structures, based on ultra thin film active and inactive components in long lengths. This effectively increases the surface area of the active plates. As a consequence, these new batteries have a higher current drain capability, lower resistance, higher energy content, lower self-discharge rate, and a wider operating temperature range than presently available solid state batteries. The preferred batteries of the present invention

are orientation tolerant, operating optimally in any configuration without regard to their vertical or horizontal positioning. Among many features and advantages, the new rechargeable batteries also provide freedom from dendrite formation, higher efficiency, lower internal resistance, greater capacity utilization, higher cycle life or cyclability, and better reliability and safety than previously available in a rechargeable metal ion battery. Moreover, these batteries are better able to tolerate overcharge and will not lead to the emission of deleterious species or outgassing. In addition, the wholly solid-state lithium-based systems of preferred embodiments of the present invention use less lithium in the cell than is typically used in existing lithium cells, thereby reducing cost as well as increasing the energy content. These and other embodiments, features and advantages of the present invention will become apparent with reference to the following description and drawings.

BRIEF DESCRIPTION OF THE DRAWINGS

For a detailed description of certain preferred embodiments of the invention, reference will now be made to the accompanying drawings wherein:

Fig. 1A shows a metallized plain polymer film with an unmetallized margin and a coating of active cathode material not on the margin.

Fig. 1B also shows a metallized plain polymer film with an unmetallized margin and layers of active anode material and solid polymer electrolyte not on the margin.

Fig. 1C shows an alternative embodiment of an electrode structure for a battery.

Fig. 2 shows the combined cathode, anode and electrolyte components of Figs. 1A and 1B and a metallized end spray.

Fig. 3 shows the layers of a thin lithium polymer electrolyte bipolar battery cell.

Fig. 4 illustrates a wound electrochemical cell for an ultra-thin battery employing layers similar to those of Figs. 1A and 1B.

DETAILED DESCRIPTION OF PREFERRED EMBODIMENTS

The inventor has developed new power source technologies that can satisfy many of the requirements which are not presently met by state-of-the-art batteries, electrochemical cells, electrodes, bipolar electrodes, and solid state electrolytes. Methods of fabricating solid state electrochemical cells, particularly high energy density cells having very thin electrode and

electrolyte structures for building ultra-thin laminar batteries are described in detail below. Also described are methods of making polymer electrolyte batteries containing such electrochemical cells, particularly rechargeable all-solid-state lithium polymer electrolyte batteries.

5 **Thin-Film Solid Polymer Electrolyte**

A very thin, low resistance, flexible film of polymer electrolyte preferably is prepared as disclosed in related application U.S. Pat. App. No. 09/388,471, filed September 2, 1999, or as described and exemplified below:

EXAMPLE 1

10 A base polymer material and a lithium salt are dissolved in a suitable liquid organic solvent (anhydrous) to form a polymer solution. Preferably the lithium salt is lithium hexafluorophosphate (LiPF_6), lithium perchlorate (LiClO_4), lithium tetrafluoroborate (LiBF_4), lithium hexafluoroarsenate (LiAsF_6), lithium tetrachloroaluminate (LiAlCl_4), lithium trifluoromethanesulfonate (LiCF_3SO_3), lithium methide ($\text{Li}(\text{SO}_2\text{CF}_3)_3$ or lithium
15 bis(trifluoromethane sulfonyl) imide ($\text{LiN}(\text{CF}_3\text{SO}_2)_2$). Preferably the base polymer material contains at least two polymers, the first being an ionically conductive polymer, the monomers of which have a hetero atom with a lone pair of electrons available for the metal ions of a metal salt to attach to and move between during conduction, when the polymer is mixed with the lithium salt. The second polymer is more inert with respect to ionic conductivity and is
20 stronger than the first polymer when they are in the form of a thin film. The conductivity is measured when each of the polymers is in the form of a thin film. Preferably the second polymer is polyester (PET), polypropylene (PP), polyethylene naphthalate (PEN), polyvinylidene fluoride (PVDF), polycarbonate (PC), polyphenylene sulfide (PPS), polytetrafluoroethylene (PTFE), or another polymer that possess stability and strength
25 characteristics similar to one of these polymers. The second polymer serves to provide the basic backbone or strength of the final polymer electrolyte when manufactured in ultra-thin films. However, if the first polymer material provides sufficient strength to the polymer electrolyte, one need not include a second polymer in the composition of the base polymer material. Likewise, if a polymer selected from the second group, in addition to meeting the
30 requisite film strength requirement also provides the desired ionic conductivity to the polymer electrolyte, a different polymer from the first group need not be included in the composition of the base polymer material. The composition of the base polymer material may comprise about 1% to 99% (by weight of said base polymer material) of the first polymer with the remainder being at least one other polymer.

A very high surface area inorganic filler – either fumed silica or alumina having an average particle size < 0.05 micron and a surface area of at least about $100 \text{ m}^2/\text{g}$ – is then extensively dispersed with a concentration in a range from about 0.1-60% by volume of the final solid polymer composition to enhance the mechanical stability and ionic conductivity of the thin film into which the polymer material with inorganic filler is cast. Further to this polymer mixture is added a finely dispersed ion conducting material having an average particle size < 0.1 micron and initial ionic conductivity of at least $2 \times 10^{-3} \text{ S/cm}$ at room temperature. Preferably the ion conducting material is a lithium ion conductor and makes up 0.1-80% by volume of the final composition or composite. The lithium ion conductor is preferably a lithium ion conducting sulfide glass or a ceramic lithium ion conductor such as lithium beta alumina or a silicate, or some other ionically conducting material of high lithium ion mobility. Care is taken to blend these constituents in an inert atmosphere since the sulfide glasses are extremely hygroscopic and may effect the performance of the final lithium batteries. This base polymer material with the plasticizer lithium salt, inorganic filler and glassy or ceramic lithium ion conductor when cast as a sheet film material and the solvent allowed to evaporate constitute the polymer electrolyte for a lithium polymer electrolyte battery. The polymer electrolyte of this composition provides a very high room temperature ionic conductivity with excellent mechanical strength.

Alternately, the base polymer material with the salt, inorganic filler, and glassy or ceramic lithium ion conductor may be intimately mixed with one another and may be cast or extruded using a hot-melt resin process. Dispersion of the inorganic filler and glassy or ceramic ion conductor into the polymer blend is performed during polymer and salt dissolution into the solvent or during extrusion of the polymer. In some embodiments, the filler is preferably 0.1-20% fumed silica or alumina having an average particle size less than or equal to 0.01 micron and the glassy or ceramic lithium ion conductor is also preferably 0.1-20% by volume having an average particle size less than or equal to 0.01 micron.

EXAMPLE 2

2 g of LiClO_4 salt was mixed with 4 g of PEO ($\text{MW } 3 \times 10^5$) and 4 g of PEG ($\text{MW } 5 \times 10^6$) in a less than 1% relative humidity dry room. The polymer/salt mixture was cast on an aluminum plate. The material was a solid gel and its ionic conductivity measured $5 \times 10^{-5} \text{ S/cm}$. The ionic conductivity of another similarly-prepared polymer/salt mixture measured $7 \times 10^{-5} \text{ S/cm}$. Knowing the base ionic conductivity of the initial polymer/salt mixture, 2 g of high surface area alumina and 5 g of finely powdered $\text{LiI-Li}_2\text{S-B}_2\text{S}_3$ glass was added to the

polymer/salt solution, and the final mixture was cast on an aluminum plate. The conductivity of the final material measured 4×10^{-4} S/cm.

EXAMPLE 3

5 2 g of lithium imide salt was mixed with a range of acrylate monomers and oligomers, including 2.1 g of isobornyl acrylate, 0.3 g of 1,6-hexanediol diacrylate, 0.8 g trimethylolpropane triacrylate, 8 g ebecryl 230 and 4 g PEO, and 0.014 g of ITX and 0.14 g of Irgacure 184 photoinitiator, dispersed in 20 volume percent isopropanol. A sample portion of this mixture was cast in a thin sheet, the IPA evaporated completely and cured using a 1000 Watts UV light. The conductivity of the cast film was measured to be 4×10^{-5} S/cm. The cast intermediate was very flexible and plastic-like. Next, 2.6 g of very high surface alumina and 8 g of finely dispersed LiI-Li₂S-B₂S₃ glass was added to the polymer/salt solution and the mixture was then cast on an aluminum plate. The solvent was completely evaporated and the polymer mixture cured using UV light. The conductivity of the final polymer electrolyte measured 6.5×10^{-4} S/cm.

EXAMPLE 4

15 mls of trimethylolpropane trimethylacrylate, 5 mls tripropylene glycol diacrylate, and 15 mls of ebecryl 230 was mixed with 3 g of lithium imide, 0.5 g Irgacure 184, 3 g of fumed silica and 5 g of the above sulfide glass. The mixture was cast between glass plates to obtain a very thin sheet and cured using UV light. The conductivity of the thin polymer sheet measured 3×10^{-4} S/cm.

A polymer electrolyte formulated as described above can be prepared in very thin form without loss of mechanical integrity, conductivity, and mechanical strength. By virtue of introducing the above-described preferred methods, which produce excellent ionic conductivity and mechanical strength of the base polymer via the two polymer hybrid design and addition of high surface area alumina or silica, and improving the ionic conductivity of the polymer electrolyte by introducing plasticizer salts and ionically conductive glassy constituents into the base polymer, as thin as 1 micron or less can be manufactured simply by selective use of the various components of the polymer electrolyte. Such polymer electrolytes are not only thin, but also truly flexible; and the thinness of the structure allows the possibility of lower resistances than are available from liquid electrolytes absorbed in traditional polypropylene separators. Traditional separator materials are usually at least 25 microns thick, and the conductivity of the solvent electrolyte is usually about 10^{-3} to 10^{-2}

S/cm at room temperature. The advantageous design of the new solid polymer electrolytes, however, makes possible thinner polymer electrolyte sections with effective resistance at least one-half of that observed in liquid electrolytes alone. For example, a liquid electrolyte with an ionic conductivity of 5×10^{-3} S/cm absorbed in a 25 micron polypropylene separator has an effective resistance of about 0.5 ohm, while a 1 micron polymer electrolyte fabricated according to one of the methods of the invention, with an ionic conductivity of 5×10^{-4} S/cm, has an effective resistance of only about 0.2 ohm. In addition, the thinness of the polymer electrolyte allows the film to be highly flexible. Although the foregoing examples are based on lithium anodes and salts, compositions and devices similar to these and based on sodium, potassium, magnesium, calcium, zinc, nickel, carbon, ion-insertion polymers, ion-insertion inorganic electrodes or tin oxide anode materials, and compatible metal salts would also be satisfactory.

The metallization material is preferably aluminum, copper, nickel, zinc, carbon, stainless steel, iconel, or any other suitable material that is compatible with the active electrode. The variation in resistivity with decreasing metal thickness is shown for an aluminum layer in Table 1.

Table 1. Resistivity as a Function of Metal Thickness for Aluminum

Resistivity (ohms per square)	Metal Thickness (Angstrom)
0.10	3000
0.30	1000
0.37	800
0.50	600
0.75	400
1.00	300
1.50	200
2.00	150
2.50	113
3.00	100

Resistivity as a function of copper thickness on metallized polyester is indicated in Table 2 below.

Table 2. Resistivity as a Function of Metal Thickness for Copper

Resistivity (ohms per square)	Metal Thickness (Angstrom)
0.1	1500
0.2	750
0.3	600
0.4	300
0.5	200
1.0	100

Each polymer substrate electrode material has different, easily identifiable, characteristics and thermal and mechanical properties, and each behaves differently depending upon its use. A suitable material can be readily selected based on the desired properties. Ideally, the thickness of the metal coating should be kept as thin as possible, while concurrently ensuring that its conductivity is very high. Preferably, the coating thickness has a conductivity of less than about 1.0 ohm per square, preferably less than 0.1 ohm per square, and more preferably about 0.01 ohm per square. This will ensure low resistance loss during current drain from the metallized substrate. The metallization may be present on only one side of the polymer layer or substrate, but is preferably provided on both sides thereof. Further, the metallization preferably is accomplished to leave an unmetallized (non-metallized) margin having a width in a range from about one millimeter (mm) to about three mm. Where the metallization is present at both sides of the polymer substrate, the non-metallization margin is provided at opposite sides of the polymer material, but on the same edge. When the substrate is coated with the active material, the coating material will be applied to the metallized portion and not the margin. The use of such substrates eliminates the use of additional carrier grids for the electrodes and instead, the metallized plastic current collector can serve the purpose of both the carrier grid for the electrodes and the battery enclosure – a major cost benefit and reducing the inactive component and making it lightweight and further increasing the energy density of the battery.

Fabrication of a Thin Film Electrode

The thin metal or metallized polymer substrate described above is coated with very thin film active anode and cathode material to complete an electrode structure that is thinner than known electrodes for thin polymer electrolyte electrochemical devices. Conventional calendaring of the anode and cathode from paste will not lead to the desired thickness.

Instead; the active anode or cathode material is milled extensively in a solvent to reduce the particle size, and then injected directly onto the substrate and subsequently drawn into thin films of any of a variety of predetermined thicknesses. Preferably the finely divided electrode material is blended with the polymer electrolyte solution.

5

EXAMPLE 5

The polymer electrolyte solution of Example 4 was mixed with 20 g of V_6O_{13} cathode active material and 0.7 g of acetylene black, and an additional 20 mls of acetonitrile, and the mixture milled for two days using zirconia grinding media. The cathode was cast directly
10 onto aluminum sheets in thin film form (20 microns), the solvent evaporated and the dried cathode/polymer mix cured using electron beam radiation.

A number of different methods may be employed to cast or coat the active electrode material/polymer electrolyte composite from a solvent-based system. Some suitable methods include knife coaters, doctor blade coaters, screen-printing, wire-wound bar coaters or Mayer
15 rods, air knife (or air doctor) coaters, squeeze roll or kiss coaters, gravure coaters, reverse roll coaters, cast film coaters, and transfer roll coaters. One coating system may be preferred over another to achieve a particular intended final result. For instance, coaters that apply an excess to the substrate web and remove the surplus coating, leaving a desired amount on the surface of the web are usually knife coaters, doctor blade coaters, bar or rod coaters, air knife
20 coaters, and squeeze roll coaters. Coaters that apply a predetermined amount of coating to the web include reverse roll coaters, gravure coaters, and transfer roll coaters. Any of these methods could be suitable for coating dry polymer electrolyte film thickness in the range of about 2 to 100 microns.

One preferred technique for applying an electrode/polymer electrolyte composite
25 material is to inject the electrode slurry with the finely dispersed inorganic and organic constituents directly onto the substrate, and then draw it out into a thin film of the desired thickness using wire wound rods (Mayer rods) with different wire diameters. The different diameter wire controls the wet slurry thickness. This method, and the other electrode deposition techniques described herein have not been used previously in the battery industry
30 to manufacture electrodes. These methods greatly facilitate obtaining an electrode material/polymer electrolyte composite film that is extremely uniform, pin-hole free, flexible and ultra-thin. Dry film thickness of 2 microns to more than 100 microns can be controlled using this process.

An alternative coating method includes incorporating each of the active electrode composites mixtures into its own aerosol mix and spraying an atomized aerosol of each material directly onto the respective opposite sides of the film substrate. This can be done rapidly while moving the substrate past the points of aerosol spray at high speed. The sprayed active electrolyte materials are cured either by drying or radiation, if the material contains radiation curable materials. Another alternative way of coating the substrate with active electrode material is to evaporate the respective electrode materials directly onto respective opposite sides of the substrate. In this case, it is preferred that the polymer electrolyte blend include a liquid radiation curable polymer capable of dissolving the metal salt and dispersing the inorganic materials. It is also preferred if inclusion of organic solvent in the electrolyte mixture can be avoided. Polymer electrolytes and active electrode/polymer electrolyte composites manufactured in this manner can result in final film thickness of less than 1 micron and more preferably in thickness of about 0.2 microns or less.

One advantage of these thinner electrode structures is that they undergo significantly less expansion and contraction during discharge and charge than is typical with other electrode structures in a battery. The polymeric structures within the battery (i.e., the metallized film substrates, and polymer film electrolyte) should also accommodate these phenomena. Since the composite cathode and anode structures must be as thin and smooth as possible, it is desirable to perform extensive, high speed, wet attrition milling of the electrode formulation. In this case, the electrode formulation preferably includes active cathode or anode material, conductive carbon, lithium salt, solvents, and a suitable binder material, so as to obtain thin electrode films similar in consistency to audio and video magnetic tapes. Final particle size of the composite electrodes is less than 0.1 micron, and preferably less than 0.05 micron. This will ensure a very thin film and smooth electrode during the coating process.

Double-Metallized Electrode Structures

As an alternative to the polymer substrate electrode structure of Fig. 1B, a preferred dual electrode structure 25 for a thin film lithium polymer electrolyte battery is fabricated by laminating anode 60 and cathode 20 elements on respective opposite sides of a double-metallized polymer substrate 13 to yield a highly flexible and strong electrode structure for the battery, as illustrated in Fig. 1C. Double-metallized polymer substrate 13 comprises polymer substrate 12 and metallization layers 16a and 16b. Non-metallized margins 14a, 14b are provided on each of the anode 60 and cathode 20 elements on the opposite sides of the metallized polymer substrate. Metal is sprayed on opposite ends of the laminated

metallized polymer substrate for the battery terminations, similar to the manner of termination shown in Fig. 2. Employing double-metallized substrates instead of singly-metallized electrode substrates further reduces the final thickness of the electrode structures, providing a ratio of substrate thickness to active electrode thickness less than about 0.5, and surface resistivity for each anode and cathode elements of less than 0.1 ohms per square. It can be readily appreciated that for metallic anode or cathode elements, the surface resistivity will be significantly lower.

After the anode and cathode composite materials have been coated or laminated onto either an ultra-thin metallic or a metallized plastic current collector, as described above, the anode and cathode elements are coated directly by a thin solid polymer electrolyte. In the double metallized half cell shown in Fig. 1C, solid polymer electrolyte layer 70 is shown overlying anode 60. Application of the solid polymer electrolyte is preferably accomplished in a single-step operation using any of the various methods of depositing the electrolytes as described in related patent application no. 09/388,741, including casting, hot-melt extrusion, hot-melt pressing, and spraying.

Bipolar Electrode Structures

Still another alternative electrode structure employing a polymer substrate is the thin film bipolar element. A bipolar structure is fabricated by laminating anode and cathode active elements on respective opposite sides of a double-metallized substrate of the same polymer material as described in the previous section entitled "Metalized Polymer Substrates" and illustrated in Fig. 1C. For a preferred bipolar structure, however, the polymer substrate is impregnated with an electronically conductive material such as carbon black or metallic elements, which are inert to the active electrode, as in the bipolar unit shown in Fig. 3. The electronically conductive material is dispersed throughout the polymeric material of the substrate. This yields a highly flexible and strong electrode structure for the battery. Preferably the electrode layers are 0.1-100 microns. For bipolar structures, the composite active cathode or anode is preferably screen-printed onto the substrate element, but one of the above-described techniques for applying active electrode composite could also be used.

Half-Element Anodes

A strong and yet flexible half-element anode consisting of a layer of anode element is made by depositing the anode material onto a very thin metallized plastic or metallic current collector. For a lithium polymer electrolyte battery, this anode half-element is lithium metal

that is preferably deposited via evaporation or sputtering technique directly onto an aluminum or nickel metallized plastic substrate or directly onto the metallic current collector (also Al or Ni) in a continuous step. Desired thickness of the plastic substrate is 0.5 to 50 microns while the desired resistivity of the metallization layer is less than 0.1 ohms per square, and the desired thickness of the lithium deposit is between 0.1 microns to 100 microns, and preferably between 0.1-10 microns. Another elemental electrode material could be substituted in place of lithium to create another type of anode battery. Lithium anodes manufactured in this manner are ultra-thin, flexible, easily handled, non-distortable, stronger and relatively inexpensive. Also, the precise capacity per unit area can be fabricated into the lithium anode composition for use in a lithium polymer electrolyte battery designed for a specific use requiring a defined battery capacity, a feature not possible with lithium metal foil anodes. This strong metallized polymer strip permits winding of the cell components in such a way that no distortion or damage to the lithium electrode occurs. Tapping the lithium using metal punching through the polymer material allows greater strength than if thin film lithium is to be used alone. If even thinner lithium metal is desired, it is preferable to evaporate it directly onto the metallized plastic current collector. The polymer material provides strength during the tapping and winding process as well as giving some backbone during the plating and stripping operation of the lithium during charge and discharge.

Very thin film battery electrode/electrolyte structures designed with very thin metallized polymer films (e.g., 1 micron) as the substrate material have several significant advantages. They make possible the fabrication of low resistance cells that can operate efficiently at the lower temperatures with significantly improved materials utilization, provide lower polarization with attendant greater energy efficiency and safety, and offer the potential for unprecedented cycle life of several tens of thousands of cycles from a bulk battery system.

Solid Polymer Electrochemical Cells and Batteries

A variety of ultra-thin solid state lithium polymer electrolyte electrochemical cells may be fabricated using the very thin solid polymer electrolytes described elsewhere herein and in U.S. Pat. App. No. 09/388,741, in combination with very thin electrode elements. Due to their flexibility and thinness, some of the more preferred electrochemical cells have conformable structures. Fabrication of a simple electrochemical cell includes laminating an anode half cell with a cathode half cell and placing a layer of solid polymer electrolyte having an initial conductivity of at least 10^{-3} to 10^{-4} S/cm between the half cells, as shown in Fig. 2.

The cell may employ the thinnest possible inactive battery components with a solid polymer electrolyte. In one such design, both sides of an ultra-thin plastic or polymer substrate are laminated with the anode and cathode elements so as to yield a preferred ultra-thin, highly flexible electrode, as shown in Fig. 3. Alternatively, multiple layers may be wound together as shown in Fig. 4. Polymer electrolytes fabricated in very thin film form are tightly rolled with thin film anode and cathode electrodes. It is preferred that once the individual thin film cell components have been laminated, the finished cell is heated to about 60-80°C for about 2 hours. This enhances the electrode/electrolyte interface and allows better interfacial adhesion and improved cyclability.

By designing an all-solid-state battery based on very thin film active and inactive components, one can effectively increase the surface area of the active plates. As a consequence, a battery will have a higher current drain capability, low resistance, higher energy content, low self-discharge, a wide operating temperature range, freedom from dendrite formation, higher efficiency, higher capacity utilization, higher cycle life, and better reliability and safety. Since there is less chance of lithium dendrite formation with the new solid polymer electrolytes, very thin polymer electrolyte membranes (e.g. 1-10 μ) as the separator and ionic carrier are feasible. Such a battery is also able to tolerate overcharge, without emission of any deleterious species or outgassing. Furthermore, when the battery design utilizes very thin metallized polymer films (e.g., about 1 micron) as the substrate material, the energy density improves by at least 100% over state-of-the-art lithium ion batteries. An especially desirable thin film lithium polymer electrolyte rechargeable battery includes an ultra thin film metallized polymer substrate about 0.5-50 microns thick. This substrate is used in conjunction with very thin film electrode and solid polymer electrolyte structures having thicknesses less than 0.5 microns. The thickness of the metallization layer on the polymer layer of the substrate is selected according to the predetermined desired conductivity of the metallization layer. The metallization layer ranges up to about 1 micron. For many applications the preferred thickness is about 0.01 micron or less.

Rechargeable Lithium Thin-Film Polymer Electrolyte Battery

EXAMPLE 6

The cathode from Example 5 was used to make a lithium bipolar battery. 10 x 10 cm sheets of the cathode were cut, a thin layer (10 microns) of the polymer electrolyte from Example 4 was cast directly onto the cathode, the solvent completely evaporated and the

cathode/pre-cured polymer electrolyte cured using UV light. The cathode/polymer electrolyte was then laminated on the back side with 1 mil thick lithium metal foil of the same area to make a bipolar plate. Two bipolar plates constructed in the above manner and two half plates of lithium and cathode were then laminated together to form a 10 V bipolar battery. Once the individual cell components have been laminated, it is preferred to heat the finished cell to just above the melting point of the polymer electrolyte. This enhances the electrode/electrolyte interface and allows better interfacial adhesion and improved cyclability. Discharge of the battery was made at 100 mA at room temperature, and pulse currents of 1 A were obtained for a duration of 10 seconds at a time.

A battery produced with all solid state components and polymer electrolytes, as exemplified above, demonstrating a room temperature ionic conductivity of at least 10^{-4} to 10^{-3} S/cm will be more stable during use and at elevated temperatures up to at least 150°C , will not outgas during overcharge, will have very low self-discharge because it is solid state, will be leakproof since there are no solvents in the system, will be very tolerant to overcharge reactions since there are no liquids to decompose, and hence will be very safe.

An ultra-thin battery according to one of the designs described herein, will better tolerate overcharge and will not lead to the emission of any deleterious species or outgassing. This is inherently provided by the fact that the polymer electrolyte is solid state. Since there are no liquid solvents within the solid polymer electrolyte, the inventor has observed that overcharge of batteries using solid polymeric electrolytes does not involve any gassing reactions. Instead the polymer merely degrades to a crystalline deposit. Hence, batteries constructed from such solid polymer electrolytes are safer than their liquid electrolyte counterparts.

In addition, the recycling of these batteries is expected to be relatively simpler than the recycling of lithium ion batteries – a major environmental benefit, especially when these batteries find use in portable electronics, telecommunications and mass electric transportation. It is also expected that the new battery technology, after more extensive testing, will prove itself to be safer than earlier technology batteries containing liquid organic solvents. Despite the fact that the polymer battery may contain lithium metal which is flammable, the fact that there are no liquid organic solvents in the battery to propagate the flammability of lithium metal and also the lithium is laminated tightly with the metallized polymer and other components of the battery will tend to limit flammability of the lithium metal. With no toxic compounds in the chemistry and no corrosive liquids in the battery (e.g.

caustic in nickel-cadmium and nickel metal hydride batteries, sulfuric acid in lead acid batteries, organic solvents in lithium ion batteries) it is expected the battery will be relatively easier to handle, recycle, and less costly than many conventional battery systems. With mass usage, it is imperative that these batteries are recycled from the beginning since lithium is a costly element and not in abundant supply, as is the case with lead, nickel, or carbon. By comparison, the lithium ion battery is only recently being discussed for recycling, after being introduced eight years ago. With mass usage of the new lithium polymer electrolyte batteries, recycling of this new technology will make a major economic and environmental difference to the world pollution problem. The inventor estimates that at least 20% of the battery cost could be recovered through recycling – a significant cost savings, especially when the market becomes mature with mass transportation applications and extensive electronics product usage.

Ultra-Thin Bipolar Batteries

An even thinner rechargeable solid polymer electrolyte battery can be fabricated with an anode half cell, a cathode half cell and at least one bipolar plate disposed therebetween. A bipolar battery has a common current collector for both the anode and the cathode and eliminates the use of one of the current collectors from each cell. For example, a lithium or cathode half cell may include a layer of the new polymer electrolyte composition with an initial ionic conductivity of at least 10^{-3} to 10^{-4} S/cm and containing at least one ionically conductive component base polymer material, a lithium salt, a high surface area inorganic filler such as alumina with an average particle size less than 0.05 microns and a surface area of at least 100 m²/g dispersed in the polymer electrolyte, and a lithium ion glassy conductor having an average particle size of less than 0.1 micron also dispersed into the polymer electrolyte. Each bipolar unit also contains on one side a layer of ultra-thin polymer electrolyte as described above. In this way, a number of these anode-current collector-cathode/polymer electrolyte bipolar units can be combined to form a multi-cell/multi-voltage battery, with the end cells being a single anode/current collector and cathode/current collector half cell at either end of the bipolar units. Bipolar half-elements are best fabricated in rectangular units using spray deposition or screen-printing methods to form the electrode coatings as well as electrolyte coatings. This kind of design is also readily adaptable for an all-solid-state design, especially a flexible polymer electrolyte battery. Batteries can be stacked in rectangular prismatic modules or the stacks can be jelly-rolled to form a cylindrical battery of higher voltage than the individual cell voltage, as described in more detail in the

following section. An ultra-thin battery assembly may be modified, if desired, by laminating the components into rolled or cylindrical geometries, or in a zig-zag or concertina configuration. Alternatively, the cylindrical geometry may be flattened after winding to form an oval shape, or flat-rectangular coupons could simply be laminated to form flat cells.

5 By reducing the lithium excesses described above for regular lithium metal cells incorporating liquid organic solvent electrolytes, and by incorporating very high surface area electrodes deposited on thin inactive substrates together with bipolar construction, the energy density of a battery can be increased to levels not previously available.

10 Ultra-thin Wound Battery

Three principal webs comprising anode 310, solid polymer electrolyte film 311 and cathode 312 are rolled or wound as illustrated in Fig. 4, to form lithium polymer electrolyte battery 315. Alternatively, the layers may be tightly stacked or laminated together. A tightly wound construction is preferred, in part because winding removes air from between the layers, and enhances creation and maintenance of continuous contact between the layers. The exercise of care in winding is essential to avoid electrical shorting of the beginning of the turns. This may be facilitated by staggering or offsetting the beginning of each layer. The conducting metallized layer deposited on the plastic film 326 may be too thin to be directly bonded to for termination purposes, so the rolled up laminate, resembling a "battery/capacitor hybrid," is taped at the edge 316, and may then be strapped in a tray (not shown) that is open on both sides. This provides access to both ends 317 and 318 of the battery 315 for scooping or, preferably, metal spraying of the roll ends, first with a high zinc content solder (harder material) followed by a regular softer "Babbitt" end spray solder (90% tin:10% zinc). The first end spray scratches the metallized surface and creates a trough to build a better electrical and mechanical contact. The tight winding and offset spacing of individual layers prevents the zinc from penetrating to the active components. This combination of end sprays also allows better contact adhesion with the final termination. The end spray also connects all the thin layers together, resulting in a continuous contact along the entire length of the film, in contrast to a cylindrical battery where only one lead contact is made at one end of the electrode. The first end spray coat (harder material) is employed in order to enhance heat transition during the scooping, to reduce shrinkage of the polymer and to increase the heat capacity on the metallized layer by providing a "heavy edge" which reinforces the metallization at the contact area.

Subsequently, aluminum leads (not shown) are soldered onto each of the ends 317 and 318 to form the final termination. The battery may then be packaged in a plastic case filled with epoxy, or conformally coated to maintain pressure on the cell, as well as to provide a hermetic seal to further protect it from humidity. This rolled laminate or battery-capacitor hybrid 315 is heated to about 80° C for a period of from 2 to 5 minutes, to improve the interface. If desired, it may be heated under vacuum before epoxying, to improve the contact between layers even further.

The anode and cathode elements are preferably laminated on both sides of a double-metallized polymer substrate material so as to yield a highly flexible electrode. By way of contrast, in a typical state-of-the-art lithium ion battery, a carbon anode is calendared on both sides of a copper current collector. In a thin film, flexible battery 315, as illustrated in Fig. 4, the lithium anode material 322 is laminated or coated in a very thin film on both sides of copper metallized polymer (e.g., PET) material 323 so that no distortion or damage to the lithium anode electrode will occur during the winding operation of the cell components. If even thinner electrode structures are desired, it is preferable that the electrode material be evaporated directly onto the metallized polymer. A similar structure (apart from the specific materials employed, e.g., aluminum rather than copper for the metallization) may exist for the cathode. An insulator sheet 325 and a metallized PET current collector 326 complete the assembly prior to winding.

In some cases, it is desirable that the coating of the anode and cathode be carried out with a margin (as shown in Figs. 1A-C and Fig. 2), especially if the tabbing or current lead can be placed at the end of the rolled structure (Fig. 4). The use of metallized plastic substrates in thickness of less than 1 micron reduces the inactive components of the cell significantly, compared to state-of-the-art lithium ion polymer electrolyte battery substrate material, thereby increasing the energy density of the battery even further. In a typical prior technology lithium metal anode battery, the ratio of substrate thickness to electrode thickness is about 0.8 to 1. In batteries assembled as described above using thinner battery elements, reduction of the ratio of substrate thickness to electrode thickness to less than 0.5, in combination with reduced electrolyte thickness and reduced dead space at the top and bottom of the jelly roll structure, significantly improves the battery's energy density, rate capability, reliability, cycle life, and safety over prior art batteries.

High-Speed Manufacturing of Ultra-Thin Polymer Electrolyte Batteries

The new, thinner battery designs lend themselves to being mass manufactured at lower cost than conventional lithium ion batteries using rapid, automated fabrication techniques. A high-speed manufacturing process for making an ultra-thin, flexible, laminar battery includes winding a cathode sheet, such as that described in Example 5 into a roll and then extruding, spraying, screen-printing, or evaporating a uniform thin layer of a solid polymer electrolyte composition described above onto the cathode sheet while the cathode roll is drawn at a uniform rate by an uptake reel. The current collector may be ultra-thin metal instead of metallized plastic. In some embodiments, it is preferred that the cathode composite sheet have narrow margins along both longitudinal edges where the plastic substrate material is uncoated by the active cathode material and the polymer electrolyte, in order to facilitate battery completion. Next the polymer electrolyte composition is cured as the composite is continuously wound by an uptake reel, to yield a roll of cathode composite.

Similarly, a roll of flexible lithium anode sheet, which is preferably lithium deposited on aluminum or nickel metallized polyester, is wound onto a reel. A uniform thin layer of the solid polymer electrolyte composition may be applied to the lithium layer similar to the manner of coating the cathode. Alternatively, a mixture of pre-cured polymer, lithium salt, filler and ionic conductor is sprayed directly onto either the anode or cathode material traveling at high speed from one roll to another. Electron beam radiation is then applied to cure the *in situ* deposited polymer electrolyte. The pre-cured polymer electrolyte mixture atomizer is preferably adjustable to various settings in order to control the exact thickness of the deposit. In this way, solid polymer electrolyte thickness of less than 0.2 microns can be achieved. Another alternative way of depositing the electrolyte coating on an electrode is to cast the electrolyte composition from a solvent based mixture and then evaporate or recover the solvent for recycling, leaving behind a thin layer of the polymer electrolyte.

Then, while simultaneously unrolling the cathode composite and the anode laminate rolls, the two sheets are placed close together and laminated together so that the adjacent layers (i.e. flexible cathode substrate/cathode/electrolyte/lithium/flexible anode substrate) are in continuous mutual contact and extraneous air is excluded from the resulting laminated cell. Another alternative way is to form the polymer electrolyte separately either through extrusion or casting, and wind the anode/polymer electrolyte/cathode in a roll-to-roll laminate form. In order to improve the interface between the layers, the laminated cell may be heated under conditions of low atmospheric pressure (vacuum). An end-spray is then applied to each end of the cylindrical cell so that tabs can be either welded or soldered to the end-spray,

maintaining pressure on the cell to keep the laminated layers in close contact during subsequent use. The end spray also protects the cell from humidity. Using a high-speed manufacturing technology like that employed today in the film capacitor making industry, electrode and electrolyte materials having the above-described compositions can be
5 manufactured in thickness of 1 micron, or even less, using an extrusion process. Alternatively, by spray-coating the polymeric materials onto the substrate, active electrode electrolyte layers can be made as thin as 0.2 micron. It is expected that the polymer electrolytes and electrodes disclosed herein will be mass produced at speeds of 500 meters per minute and in thickness of 0.2 to 100 microns for producing thin film rechargeable
10 lithium polymer electrolyte batteries. A similar high-speed production process, using a continuous roll-to-roll winding of the various battery elements in the form of strong, thin sheets will reduce the cost of manufacturing laminar solid polymer electrolyte batteries significantly. Another important advantage of this high-speed method of production is that batteries are produced which contain more uniformly thick, pin-hole free electrolyte layers
15 than is possible with conventionally produced polymer electrolyte batteries.

While the preferred embodiments of the invention have been shown and described, modifications thereof can be made by one skilled in the art without departing from the spirit and teachings of the invention. The embodiments described herein are exemplary only, and are not intended to be limiting. For instance, while certain preferred lithium solid polymer electrolytes,
20 and lithium polymer electrolyte batteries, electrodes, and components thereof are emphasized in the description and examples, the methods and compositions of the invention can be readily extended to any battery chemistry, both primary and secondary to provide other superior performance batteries. Many variations and modifications of the invention disclosed herein are possible and are within the scope of the invention. Accordingly, the scope of protection is not
25 limited by the description set out above, but is only limited by the claims which follow, that scope including all equivalents of the subject matter of the claims.

CLAIMS

What is claimed is:

1. An all-solid-state laminar electrochemical cell for a battery comprising:

an anode layer about 0.1-100 μ thick;
a cathode layer about 0.1-100 μ thick;
an anode current collector about 0.5-50 μ thick attached to said anode;
a cathode current collector about 0.5-50 μ thick attached to said cathode; and
a layer of solid polymer electrolyte disposed between said anode and cathode layers.

2. An all-solid-state laminar electrochemical cell for a battery comprising:

an anode layer;
a cathode layer;
an anode current collector attached to said anode;
a cathode current collector attached to said cathode; and
a layer of solid polymer electrolyte disposed between said anode and cathode layers, said solid polymer electrolyte being a cationic conductor having a conductivity of at least 10^{-4} S/cm at 25°C and comprising a mixture of

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g; and

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

3. The all-solid-state laminar electrochemical cell of claim 2 wherein said anode layer is about 0.1-100 μ thick, said cathode layer is about 0.1-100 μ thick, said anode and cathode current collectors are each 0.5-50 μ thick and said solid polymer electrolyte is about 0.2-100 μ thick.

4. The all-solid-state laminar electrochemical cell of claim 2 wherein said anode is lithium, said metal salt is a lithium salt, and said ionic conducting material is a lithium ion conducting material.

5. The all-solid-state laminar electrochemical cell of claim 4 wherein said lithium salt is chosen from the group consisting of lithium hexafluorophosphate (LiPF_6), lithium perchlorate (LiClO_4), lithium tetrafluoroborate (LiBF_4), lithium hexafluoroarsenate (LiAsF_6), lithium tetrachloroaluminate (LiAlCl_4), lithium trifluoromethanesulfonate (LiCF_3SO_3), lithium methide ($\text{LiC}(\text{SO}_2\text{CF}_3)_3$) and lithium bis(trifluoromethane sulfonyl) imide ($\text{LiN}(\text{CF}_3\text{SO}_2)_2$).

6. The all-solid-state laminar electrochemical cell of claim 2 wherein said base polymer material comprises at least two polymers, the first said polymer being an ionically conductive polymer, the monomers of which have a hetero atom with a lone pair of electrons available for the metal ions of a metal salt to attach to and move between during conduction, when said first polymer is mixed with a metal salt.

7. The all-solid-state laminar electrochemical cell of claim 6 wherein the composition of said base polymer material comprises about 1% to 99% (by weight of said base polymer material) of one polymer and the remainder of said base polymer material comprises at least one other polymer.

8. The all-solid-state laminar electrochemical cell of claim 6 wherein said first polymer is chosen from the group consisting of linear polymers, random copolymers, block copolymers, comb-branched block copolymers, network structures, single ion conductors, polyvinylidene fluoride or chloride and copolymers of their derivatives, poly(chlorotrifluoroethylene), poly(ethylene-chlorotrifluoro-ethylene), and poly(fluorinated ethylene-propylene).

9. The all-solid-state laminar electrochemical cell of claim 8 wherein said first polymer is chosen from the group consisting of polyethylene oxide (PEO), oxymethylene linked PEO, PEO-PPO-PEO crosslinked with trifunctional urethane; poly(bis(methoxy-ethoxy-ethoxide))-phosphazene (MEEP); a triol-type PEO crosslinked with difunctional urethane, poly((oligo)oxyethylene)methacrylate-co-

alkali metal methacrylate; polyacrylonitrile (PAN), polymethylmethacrylate (PMMA), polymethylacrylonitrile (PMAN); polysiloxanes and their copolymers and derivatives, polyvinylidene fluoride or chloride and copolymers of their derivatives, poly(chlorotrifluoro-ethylene), poly(ethylene-chlorotrifluoroethylene),
5 poly(fluorinated ethylene-propylene), acrylate-based polymer, other similar solvent-free polymers, combinations of the foregoing polymers either condensed or crosslinked to form a different polymer, and physical mixtures of any of the foregoing polymers,

provided said first polymer is combinable with a lithium salt such that the
10 ionic conductivity of said first polymer is enhanced.

10. The all-solid-state laminar electrochemical cell of claim 8 wherein said second of
said at least two polymers is more inert with respect to ionic conductivity and is
stronger than said first polymer when each said polymer is in the form of a thin film.

11. The all-solid-state laminar electrochemical cell of claim 10 wherein said second
polymer is chosen from the group consisting of polyester (PET), polypropylene (PP),
polyethylene naphthalate (PEN), polyvinylidene fluoride (PVDF), polycarbonate (PC),
polyphenylene sulfide (PPS), polytetrafluoroethylene (PTFE), and other polymer
20 materials that possess stability and strength characteristics similar to one of said group
of polymers.

12. The all-solid-state laminar electrochemical cell of claim 2 wherein said inorganic
filler is chosen from the group consisting of fumed silica and alumina.

13. The all-solid-state laminar electrochemical cell of claim 2 wherein said ionic
conducting material is chosen from the group consisting of glassy lithium ion
conductors and ceramic lithium ion conductors.

14. The all-solid-state laminar electrochemical cell of claim 3 in a stacked configuration.

15. The all-solid-state laminar electrochemical cell of claim 3 having a conformable
structure.

16. The all-solid-state laminar electrochemical cell of claim 3 in a wound configuration.
17. The all-solid-state laminar electrochemical cell of claim 3 wherein each said current collector is a very thin structure chosen from the group consisting of metallic elements less than 10 μ thick and metallized plastics 0.5 to 50 μ thick.
18. An elemental anode for a polymer electrolyte battery comprising a 0.1-50 μ thick layer of elemental anode material overlying a current collector chosen from the group consisting of:
- metallic current collectors about 1-10 μ thick, and
- metallized polymer current collectors comprising a polymer substrate about 0.5-50 μ thick and a metallization layer up to about 1 μ thick overlying said polymer substrate.
19. The elemental anode of claim 18 wherein said metallic current collector and said metallization layer are selected from the group consisting of aluminum, nickel and inconel, copper, stainless steel, zinc, carbon and gold.
20. The elemental anode of claim 1 wherein said current collector is a metallic current collector about 2-5 μ in thickness.
21. The elemental anode of claim 1 wherein said metallic current collector is less than 2 μ thick.
22. The elemental anode of claim 18 wherein said current collector is a metallized polymer current collector comprising a polymer substrate about 0.5-50 μ thick and a metallization layer up to 1 μ thick overlying at least a portion of said polymer substrate.
23. The elemental anode of claim 22 wherein said metallized polymer current collector includes a non-metallized margin on said polymer substrate.

24. The elemental anode of claim 1 wherein said anode material is chosen from the group consisting of lithium, sodium, potassium, magnesium, calcium, zinc and nickel.

25. An anode half-element for a polymer electrolyte battery comprising the elemental
5 anode of claim 18 and a layer of solid polymer electrolyte overlying said layer of elemental anode material, said solid polymer electrolyte is a cationic conductor having a conductivity greater than 1×10^{-4} S/cm at 25°C or below and comprising a mixture of:

10 a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$; and

15 an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

26. An anode half-element for a lithium polymer electrolyte battery comprising a
substrate about $0.5\text{-}50\mu$ thick containing a current collector;

20 a thin layer of lithium anode material about $1\text{-}100\mu$ thick overlying said current collector, said lithium anode material comprises a mixture of:

an active anode substance;

25 a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt;

said lithium salt;

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$; and

30 a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

27. The anode half element of claim 26 wherein said active anode substance is chosen from the group consisting of ion-insertion polymers, ion-insertion inorganic electrodes, carbon insertion electrodes and tin oxide.

28. A cathode half element for a lithium polymer electrolyte battery comprising a 0.1-100 μ thick layer of cathode material overlying a current collector chosen from the group consisting of:

metallic current collectors about 1-10 μ thick, and

metallized polymer current collectors about 0.5-50 μ thick comprising a polymer substrate and a metallization layer up to about 1 μ thick.

29. The cathode half element of claim 28 wherein said cathode material comprises a mixture of

an active cathode substance;

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g; and

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

30. The cathode half element of claim 29 wherein said active cathode substance is chosen from the group consisting of oxides, sulfides and selenides.

31. The cathode half element of claim 30 wherein said active cathode substance is chosen from the group consisting of MnO₂, LiMn₂O₄, Li_xMnO₂, MoS₂, MoS₃, MoV₂O₈, CoO₂, Li_xCoO₂, V₆O₁₃, V₂O₅, V₃O₈, VO₂, V₂S₅, TiS₂, NbSe₃, Cr₂O₅, Cr₃O₈, WO₃, Li_xNiO₂, Li_xNi_yCo_zO₂, Li_xNi_yMn_zO₂ and Li_xCo_yMn_zO₂.

32. The cathode half element of claim 29 wherein said active cathode substance is a lithium doped electronically conducting polymer.

33. The cathode half element of claim 32 wherein said electronically conducting polymer is chosen from the group consisting of polypyrrole, polyaniline and polyacetylene.
- 5 34. A thin electrode for an all-solid-state lithium polymer electrolyte battery comprising a substrate about 0.5-50 μ thick, current collector, and a layer of active electrode/polymer electrolyte composite about 0.1-100 μ thick overlying a side of said substrate, said active electrode/polymer electrolyte composite comprising a mixture of
- 10 an active electrode material;
a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt;
a lithium salt;
an inorganic filler having an average particle size < 0.05 micron in diameter
- 15 and a surface area of at least about 100 m²/g; and
an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.
- 20 35. The thin electrode of claim 34 wherein said substrate is an ultra-thin film metallized polymer substrate comprising a polymer layer about 0.5-50 microns thick and a metallization layer up to 1 μ thick and having a conductivity of about 0.01 - 1.0 ohm per square, said metallization layer overlying a side of said polymer layer and constituting said current collector.
- 25 36. The thin electrode of claim 35 wherein said metallization layer is $\leq 0.01\mu$ thick.
37. A bipolar electrode for a battery comprising:
a flexible polymer substrate about 0.5-50 μ thick and having two opposite sides;
- 30 a 0.1-100 μ thick anode layer overlying one said side;
a 0.1-100 μ thick cathode layer overlying the other said side, the ratio of substrate thickness to anode or cathode layer thickness being less than about 0.5, and

the surface resistivity for each said anode and cathode layer being less than 0.1 ohms per square.

38. The bipolar electrode of claim 37 wherein said flexible polymer substrate is impregnated with an electronically conductive element that is chemically inert to said anode and cathode layers.

39. The bipolar electrode of claim 37 further comprising a first metallization layer up to 1 μ thick disposed between said anode layer and said one side, and a second metallization layer up to 1 μ thick disposed between said cathode layer and said other side.

40. The bipolar electrode of claim 37 further comprising a layer of solid polymer electrolyte overlying at least one of said anode layer and said cathode layer, said solid polymer electrolyte comprising a mixture of

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g; and

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

41. The bipolar electrode of claim 38 wherein said electronically conductive element has a conductivity greater than 10^2 S/cm at 25°C and is chosen from the group consisting of conductive carbon, an electronically conductive polymer and finely divided metal.

42. The bipolar electrode of claim 41 wherein said electronically conductive polymer is chosen from the group consisting of polyacetylene, polypyrrole and polyaniline.

43. A defined capacity battery comprising the electrochemical cell of claim 2 wherein said anode layer comprises a predetermined quantity of an element chosen from the

group consisting of lithium, sodium, potassium, magnesium, calcium, zinc, carbon and nickel per unit area of said anode layer.

44. A leakproof battery comprising the all-solid-state laminar electrochemical cell of claim 2 in a moisture resistant covering.

45. An overcharge tolerant battery comprising the all-solid-state laminar electrochemical cell of claim 2.

46. The overcharge tolerant battery of claim 45 further comprising an overcharge indicator.

47. An orientation tolerant polymer electrolyte battery having an all-solid-state composition comprising the laminar electrochemical cell of claim 2.

48. An all-solid-state battery comprising the wound electrochemical cell of claim 16 wherein said cell comprises a cylindrical configuration having cylinder ends with metallic terminations.

49. The all-solid-state battery of claim 48 wherein said terminations comprise at least one end coating on said cylinder ends.

50. The all-solid-state battery of claim 49 further comprising a moisture-resistant covering.

51. The all-solid-state battery of claim 2 wherein said cell comprises a concertina configuration.

52. An all-solid-state lithium polymer electrolyte battery comprising:

a pair of spaced-apart flexible thin film electrodes, each said electrode comprising a layer of active electrode material overlying a polymer substrate having an adherent electrically conductive layer disposed thereon and/or having an electrically conductive material dispersed in said polymer substrate; and

a resilient and flexible thin film solid polymer electrolyte tightly disposed between said pair of electrodes, said electrolyte having a conductivity of at least about 1×10^{-4} S/cm at 25°C and containing a mixture of:

at least two polymers, one of which is an ionically conductive polymer having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt,

said lithium salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$, and

a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

53. The battery of claim 52 wherein said inactive polymer substrate of each said electrode comprises a polymer material selected from the group of polymers consisting of PET, PP, PPS, PEN, PVDF and PE.

54. The battery of claim 52 wherein said inactive polymer substrate is a metallized polymer substrate having a thin metallization layer as said adherent electrically conductive layer.

55. The battery of claim 54 wherein said metallized polymer substrate is about 0.5 to 50 microns thick and said thin metallization layer has a resistivity of about 0.01 to 1 ohm per square.

56. The battery of claim 54 wherein said metallized polymer substrate further comprises a first side containing a first non-metallized margin.

57. The battery of claim 56 wherein said non-metallized margin extends about 1 to 3 mm from an edge of said first side to an edge of said first metallization layer.

58. The battery of claim 57 further comprising a second metallization layer adhered to a second side of said polymer layer, said second side of said polymer layer comprising

a second unmetallized margin extending about 1 to 3 mm from said edge to an edge of said second metallization layer.

59. In a lithium polymer electrolyte battery containing a lithium metal anode layer, a cathode layer, an anode current collector attached to said anode layer, a cathode current collector attached to said cathode layer, and a layer of lithium polymer electrolyte disposed between said anode and cathode layers, the improvement comprising substitution of a lithium solid polymer electrolyte having a conductivity of at least about 10^{-3} to 10^{-4} S/cm at 25°C for said at least one polymer electrolyte layer, said lithium solid polymer electrolyte comprising a mixture of

a polymer or polymer blend having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt,

said lithium salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$, and

a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C,

the improved battery having a cell resistance less than or equal to that of a liquid electrolyte lithium battery.

60. The improved lithium polymer battery of claim 59 wherein the improvement further comprises substitution for at least one said current collector an ultra-thin current collector chosen from the group consisting of metallic elements less than 10μ thick and metallized plastics 0.5 to 50μ thick.

61. The improved lithium polymer battery of claim 60 wherein the improvement further comprises substituting for at least one of said anode and cathode layers a flexible ultra thin film electrode containing a metallized polymer substrate having an inactive polymer layer about 0.5 - 50 microns thick, a metallization layer up to about 0.01 micron thick overlying said inactive polymer layer, and a layer of active electrode material up to about 5 microns thick overlying said metallization layer; and substituting for said

polymeric electrolyte a lithium solid polymer electrolyte in the form of a thin film up to about 5 microns thick, said lithium solid polymer electrolyte comprising a mixture of

a polymer or polymer blend having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt,

said lithium salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$, and

a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

62. The battery of claim 61 wherein said filler is 0.1-20% (by volume of solid polymer electrolyte) high surface area filler having an average particle size ≤ 0.01 micron and chosen from the group consisting of fumed silica and alumina.

63. The battery of claim 61 wherein said lithium ion conducting material is 0.1-80% sulfide glass (by volume of solid polymer electrolyte).

64. The battery of claim 63 wherein said lithium ion conducting material is a ceramic ion conductor chosen from the group consisting of lithium beta alumina and silicates.

65. The battery of claim 59 wherein said cathode layer comprises a solid polymeric composite containing a mixture of

active cathode material having an average particle size ≤ 0.1 micron; and

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$; and

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

66. A rechargeable all-solid-state lithium polymer electrolyte battery comprising:

an ultra thin lithium anode chosen from the group consisting of a metallic lithium element about 0.1 to 100 microns thick and a lithium metal layer about 0.1 to 100 microns thick over the metallization layer of a metallized polymer substrate, said metallized polymer substrate having an inactive polymer layer about 0.5 to 50 microns thick and a metallization layer up to 0.01μ thick overlying said inactive polymer layer;

an ultra thin film cathode layer containing a metallized polymer substrate said metallized polymer substrate having an inactive polymer layer about 0.5 to 50 microns thick and a metallization layer up to about 0.01μ thick overlying said inactive polymer layer, and a layer of active cathode material $0.1\text{--}100\mu$ thick overlying said metallization layer; and

a polymer electrolyte layer up to 5μ thick disposed between said anode and cathode layers, said lithium polymer electrolyte having a conductivity greater than 1×10^{-4} S/cm at 25°C or below and comprising a mixture of

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt,

said lithium salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$, and

a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C .

67. The rechargeable all-solid-state lithium polymer electrolyte battery of claim 66 wherein said polymer electrolyte comprises:

about 30 to 95% (by weight of solid polymer electrolyte) base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt;

about 1 to 25% (by weight of solid polymer electrolyte) said lithium salt;

* about 0.1-80% (by volume of solid polymer electrolyte) inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$; and

about 0.1-60% (by volume of solid polymer electrolyte) ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least $2 \times 10^{-3} \text{ S/cm}$ at 25°C .

68. The rechargeable all-solid-state lithium polymer electrolyte battery of claim 67 wherein the concentration of said inorganic filler is about 0.1-20% (by volume of solid polymer electrolyte) and the concentration of said ionic conducting material is about 0.1-20% (by volume of solid polymer electrolyte).

69. A method of making an ultra-thin solid polymer electrochemical cell comprising:

preparing an anode sheet by applying active anode material to an inactive anode substrate such that the ratio of anode substrate thickness to anode sheet thickness is ≤ 0.8 and the conductivity is ≤ 0.1 ohms per square;

preparing a cathode sheet by applying active cathode material to an inactive cathode substrate such that the ratio of cathode substrate thickness to cathode sheet thickness is ≤ 0.8 and the conductivity is ≤ 0.1 ohms per square;

preparing a solid polymer electrolyte having a conductivity of at least 10^{-4} S/cm at 25°C and comprising a mixture of

a polymer or polymer blend having an initial conductivity of at least about $1 \times 10^{-5} \text{ S/cm}$ at 25°C when combined with a lithium salt,

said lithium salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$,

a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least $2 \times 10^{-3} \text{ S/cm}$ at 25°C ,

an organic solvent, and

a curing agent;

coating said cathode sheet with said solid polymer electrolyte, evaporating said solvent, and curing said polymer electrolyte to yield an electrolyte-coated cathode sheet;

laminating said electrolyte-coated cathode sheet and said anode sheet together such that said electrolyte coat is between said cathode active material and said anode active material, whereby a cathode/electrolyte/anode laminate is formed;

optionally winding said laminate; and

optionally shaping said laminate.

70. The method of claim 69 wherein said step of preparing an anode sheet comprises applying active anode material to an inactive anode substrate containing an inactive polymer substrate and a thin metallization layer overlying at least a portion of said polymer substrate, such that the ratio of anode substrate thickness to anode sheet thickness is ≤ 0.8 and the conductivity is ≤ 0.1 ohms per square; and

said step of preparing a cathode sheet comprises applying active cathode material to an inactive cathode substrate containing an inactive polymer substrate and a thin metallization layer overlying at least a portion of said polymer substrate, such that the ratio of cathode substrate thickness to cathode sheet thickness is ≤ 0.8 and the conductivity is ≤ 0.1 ohms per square.

71. A method of making a strong and flexible anode/electrolyte half-element for a laminar battery comprising:

depositing a 0.1-100 microns thick layer of elemental metal onto an electrically conductive substrate chosen from the group consisting of

metallic current collectors about 1-10 μ thick, and

metallized polymer current collectors comprising a polymer substrate about 0.5-50 μ thick and a metallization layer up to about 1 μ thick overlying said polymer substrate, said metallization layer having a resistivity less than 1 ohm per square, said elemental metal chosen from the group consisting of lithium, sodium, potassium, magnesium, calcium, zinc, carbon and nickel; and

applying a layer of solid polymer electrolyte onto said layer of elemental metal, said solid polymer electrolyte having a conductivity greater than 1×10^{-4} S/cm at 25°C or below and comprising a mixture of:

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

5 an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$; and

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C.

10

72. The method of claim 71 wherein said depositing comprises evaporating or sputtering said elemental metal directly onto said current collector in a continuous step.

15

73. The method of claim 72 wherein said depositing comprises applying a predetermined quantity of elemental metal to yield a precise elemental metal capacity per unit area of anode when said anode half-element is employed in a battery.

74. The method of claim 73 further comprising tabbing said elemental metal layer by metal punching through said polymer substrate.

20

75. A method of making a strong and flexible electrode/electrolyte half-element for a laminar battery, the method comprising:

depositing a 0.1-100 microns thick layer of active electrode material onto an electrically conductive substrate chosen from the group consisting of

25

metallic current collectors about 1-10 μ thick, and

metallized polymer current collectors comprising a polymer substrate about 0.5-50 μ thick and a metallization layer up to about 1 μ thick overlying said polymer substrate, said metallization layer having a resistivity less than 1 ohm per square;

30

applying a layer of solid polymer electrolyte onto said layer of active electrode material, said solid polymer electrolyte having a conductivity greater than 1×10^{-4} S/cm at 25°C or below and comprising a mixture of:

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt;

a metal salt;

5 an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$;

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C;

10 optionally, an inorganic solvent; and
optionally, a curing agent.

76. The method of claim 75 wherein said step of depositing active electrode material and said step of applying said solid polymer electrolyte include casting each of said active
15 electrode material and said solid polymer electrolyte using an apparatus independently chosen from the group consisting of knife coaters, doctor blade coaters, wire-wound bar coaters, air knife coaters, squeeze roll, gravure coaters, reverse roll coaters, cast film coaters and transfer roll coaters.

20 77. The method of claim 75 wherein said step of depositing active electrode material and said step of applying said solid polymer electrolyte include extruding each of said active electrode material and said solid polymer electrolyte.

78. A method of making a thin bipolar battery comprising laminating tightly together at least
25 one layer of bipolar unit between a layer of first anode and a layer of first cathode to provide a stack having laminar ends,

said bipolar unit comprising

a flexible first polymer substrate about $0.5\text{-}50\mu$ thick and having two opposite sides and two opposite edges, said first polymer substrate including
30 an electrically conductive material dispersed therein;

a $0.1\text{-}100\mu$ thick second anode layer overlying one said side;

a $0.1\text{-}100\mu$ thick second cathode layer overlying the other said side,
the ratio of substrate thickness to second anode or second cathode layer

thickness being less than about 0.5, and the surface resistivity for each said second anode and second cathode layer being less than 0.1 ohms per square; optionally, a metallization layer up to 1 μ thick disposed between said anode layer and said one side, and a metallization layer up to 1 μ thick disposed between said cathode layer and said other side; and

a layer of solid polymer electrolyte overlying at least one of said second anode layer and said second cathode layer, said solid polymer electrolyte having a conductivity of at least 10^{-4} S/cm at 20°C and comprising a mixture of

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt,

a metal salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g, and

an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C;

said first anode comprising

a 0.1-100 μ thick layer of elemental anode material overlying a current collector chosen from the group consisting of:

metallic current collectors about 1-10 μ thick, and

metallized polymer current collectors comprising a second polymer substrate about 0.5-100 μ thick and a metallization layer up to about 1 μ thick overlying said second substrate, and

a layer of said solid polymer electrolyte overlying said layer of elemental anode material, and
said first cathode comprising:

a 0.1-100 μ thick layer of cathode material overlying a current collector chosen from the group consisting of:

metallic current collectors about 1-10 μ thick, and

metallized polymer current collectors about 0.5-50 μ thick comprising a polymer substrate and a metallization layer up to about 1 μ thick,

orienting said first anode and first cathode such that said first anode elemental material opposes said active cathode substance of said second cathode of a bipolar unit, with a layer of solid polymer electrolyte disposed therebetween; and applying terminations to said first anode and first cathode.

79. A high-speed manufacturing method for producing an ultra-thin laminar battery, the method comprising:

winding a 0.1-100 μ thick cathode sheet into a roll, said cathode sheet comprising

a flexible inactive polymer substrate and a layer of active cathode material overlying said substrate, said substrate optionally including an uncoated margin along a longitudinal edge of said substrate;

extruding a uniform layer of a lithium solid polymer electrolyte composition onto said cathode sheet while said roll is drawn at a uniform rate by an uptake reel, to obtain a roll of flexible substrate/active cathode/electrolyte composite, said lithium solid polymer electrolyte comprising a mixture of

a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a lithium salt,

said lithium salt,

an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g, and

a lithium ion conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C;

curing said extruded polymer electrolyte composition as said composite is continuously wound by said uptake reel, to yield a roll of cathode/electrolyte laminate;

winding a lithium anode sheet to provide a roll of anode laminate, said lithium anode sheet comprising a 0.1-100 μ thick layer of elemental anode material overlying a current collector chosen from the group consisting of:

metallic current collectors about 1-10 μ thick, and

metallized polymer current collectors comprising a polymer substrate about 0.5-50 μ thick and a metallization layer up to about 1 μ thick overlying said polymer substrate;

winding a 0.5-50 μ thick sheet of inert plastic to provide a first roll of plastic;

winding a 0.5-50 μ thick sheet of inert plastic to provide a second roll of plastic; while simultaneously unrolling said rolls of cathode/electrolyte laminate, anode laminate and first and second rolls of plastic, laminating said cathode/electrolyte laminate and said anode laminate together, between layers of said plastic sheet, whereby adjacent layers of plastic sheet/cathode/electrolyte/lithium anode/plastic sheet are in continuous mutual contact and extraneous air is excluded from between the layers, whereby a cylindrical laminated cell having first and second electrode edges is obtained;

applying a current collector to each said edge;

attaching leads to the respective current collectors;

maintaining pressure on said laminated cell;

optionally heating said laminated cell;

optionally applying a vacuum to said laminated cell during said heating; and hermetically sealing said cell.

80. The method of claim 79 wherein said extruding step comprises extruding a uniform thin layer of said solid polymer electrolyte composition onto said cathode sheet in such a manner as to leave an uncovered margin on said substrate along a longitudinal edge of said substrate.

81. The method of claim 79 further comprising heating said laminated cell and, optionally, applying a vacuum during said heating.

82. A method of making a thin film electrode/electrolyte half-element for a battery comprising:

laminating together a sheet of thin film composite electrode and a sheet of thin film solid polymer electrolyte, to form a laminate;

optionally heating said laminate; and

optionally, applying a vacuum to said laminate during said heating, said composite electrode comprising a 0.1-100 μ thick polymeric composite layer overlying a 0.5-50 μ thick flexible plastic substrate, said polymeric composite layer comprising a mixture of active electrode material having an average particle size ≤ 0.1 micron, a first base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with a metal salt, said metal salt, an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g, and an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C; and

said solid polymer electrolyte having a conductivity greater than 1×10^{-4} S/cm at 25°C or below and comprising a mixture of a second base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25°C when combined with said metal salt, said metal salt, an inorganic filler having an average particle size < 0.05 micron in diameter and a surface area of at least about 100 m²/g, and an ionic conducting material having an average particle size < 0.1 micron in diameter and an initial ionic conductivity of at least 2×10^{-3} S/cm at 25°C, said first and second base polymer materials being the same or different.

83. The method of claim 82 wherein said laminating step includes roll-to-roll winding of said thin film composite electrode sheet and said thin film solid polymer electrolyte sheet.

84. The method of claim 83 wherein said solid polymer electrolyte sheet comprises a pre-cured, extruded polymer electrolyte composition.

1/3

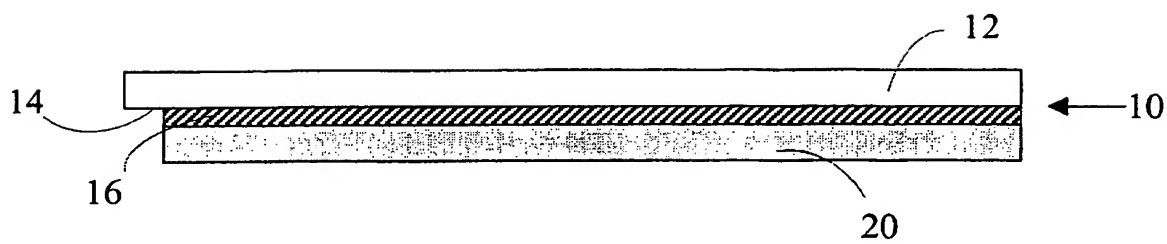


FIG. 1A

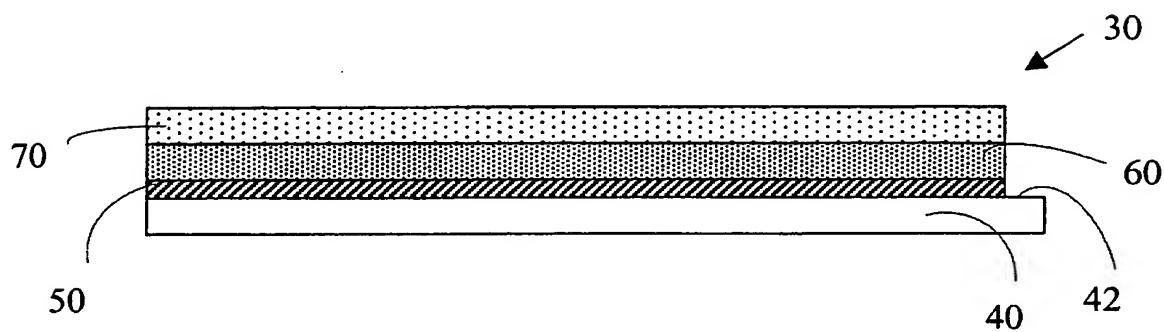


FIG. 1B

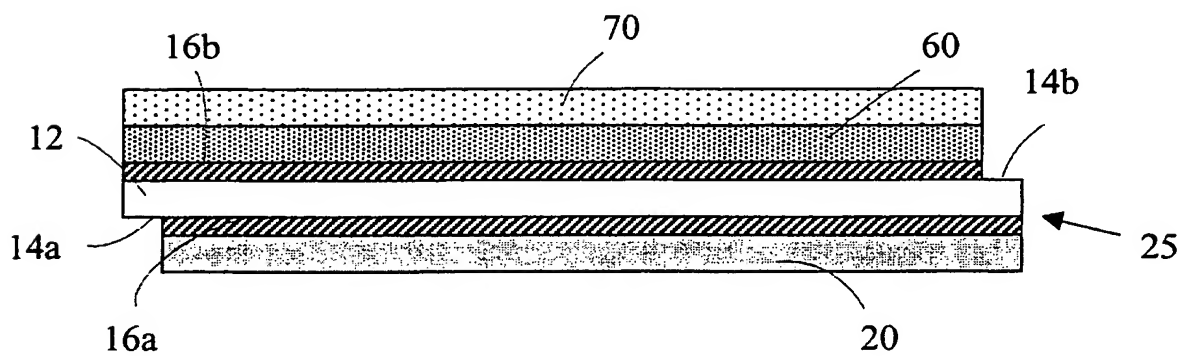


FIG. 1C

2/3

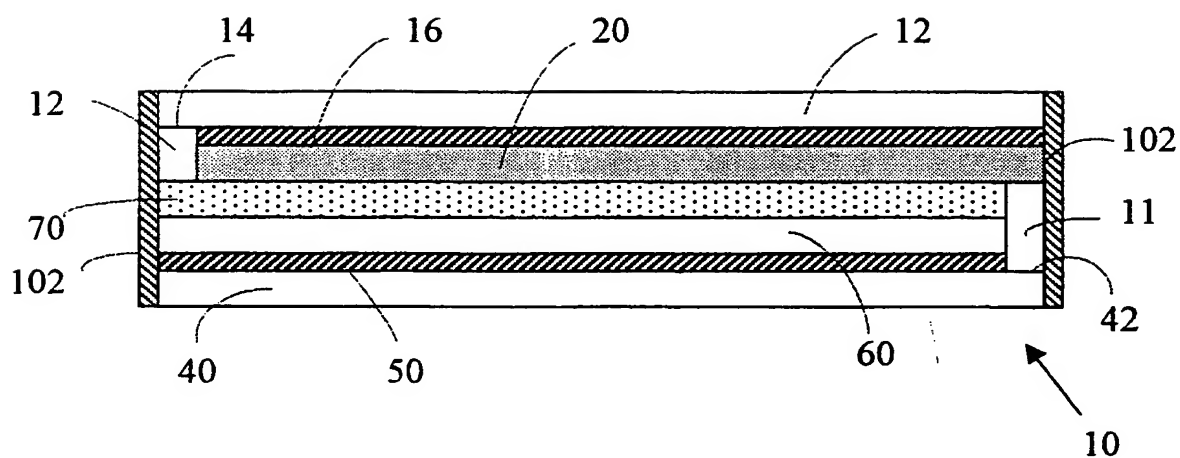


FIG. 2

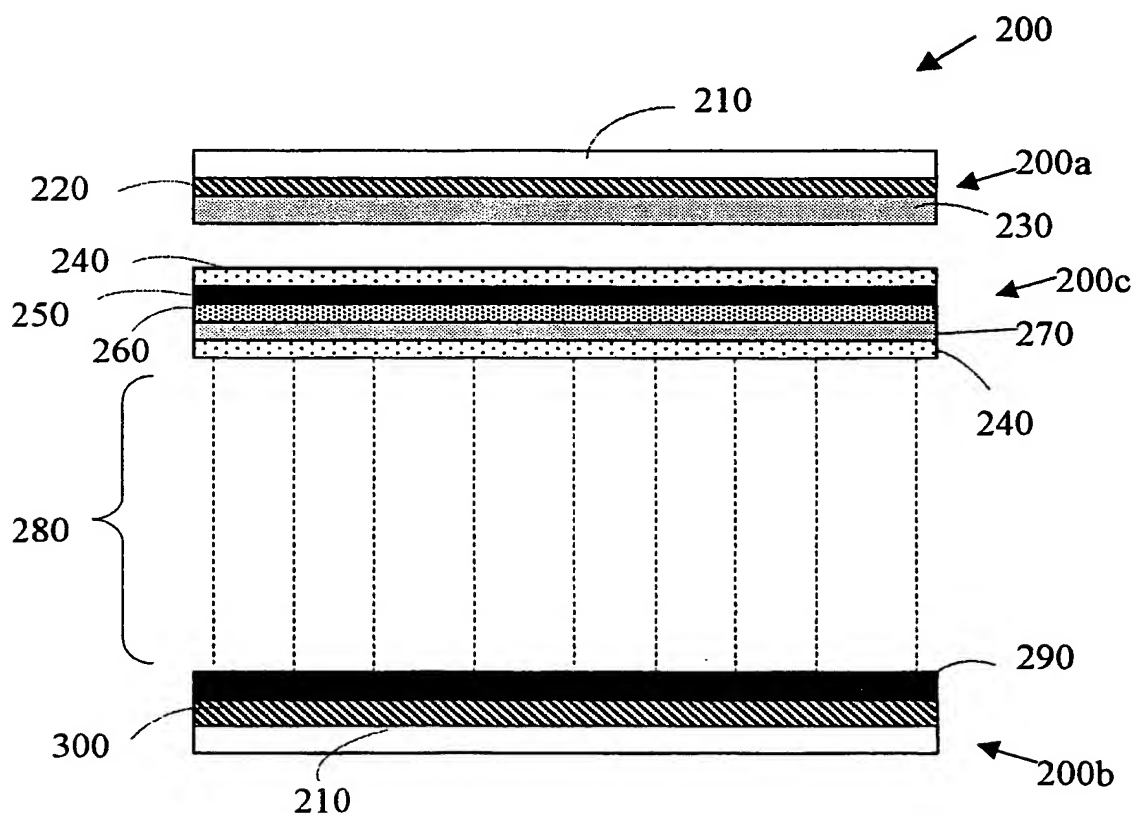


FIG. 3

3/3

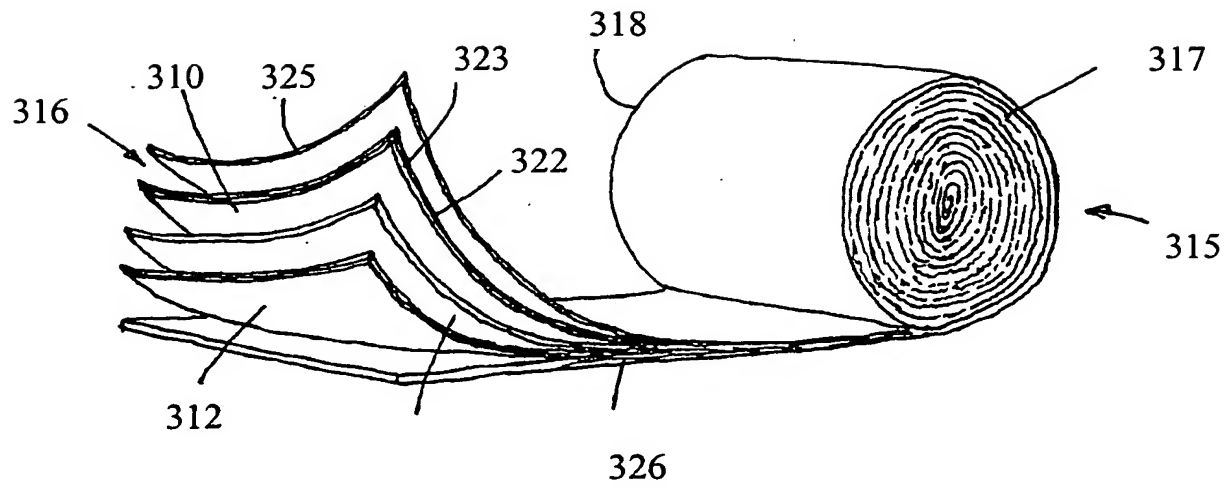


FIG. 4

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
8 March 2001 (08.03.2001)

PCT

(10) International Publication Number
WO 01/17052 A3

(51) International Patent Classification⁷: **H01M 10/40.**
H01G 9/02, H01B 1/12

(21) International Application Number: PCT/US00/22917

(22) International Filing Date: 21 August 2000 (21.08.2000)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
09/388,733 2 September 1999 (02.09.1999) US

(71) Applicant: **LITHIUM POWER TECHNOLOGIES, INC.** [US/US]; 20955 Morris Avenue, Manvel, TX 77578 (US).

(72) Inventor: **MUNSHI, M., Zafar, A.**; P.O. Box 978, Manvel, TX 77578 (US).

(74) Agents: **MINTZ, Carol, E. et al.**; Conley, Rose & Tayon, P.C., P.O. Box 3267, Houston, TX 77253-3267 (US).

(81) Designated States (*national*): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CR, CU, CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, UZ, VN, YU, ZA, ZW.

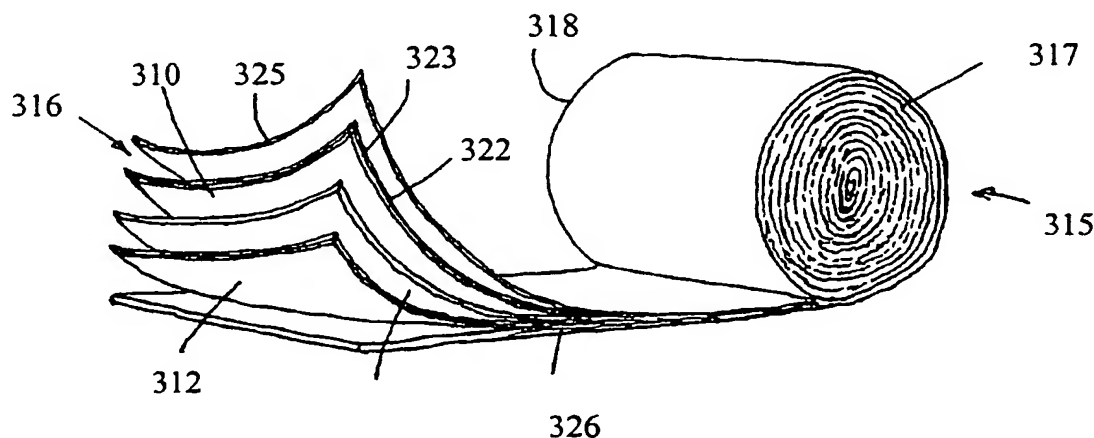
(84) Designated States (*regional*): ARIPO patent (GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG).

Published:
— with international search report

(88) Date of publication of the international search report:
21 February 2002

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: ALL-SOLID-STATE ELECTROCHEMICAL DEVICE AND METHOD OF MANUFACTURING



(57) Abstract: All-solid-state electrochemical cells and batteries employing very thin film, highly conductive polymeric electrolyte and very thin electrode structures are disclosed, along with economical and high-speed methods of manufacturing. A preferred embodiment is a rechargeable lithium polymer electrolyte battery. New polymeric electrolytes employed in the devices are strong yet flexible, dry and non-tacky. The new, thinner electrode structures have strength and flexibility characteristics very much like thin film capacitor dielectric material that can be tightly wound in the making of a capacitor. A wide range of polymers, or polymer blends, characterized by high ionic conductivity at room temperature, and below, are used as the polymer base material for making the solid polymer electrolytes. The preferred polymeric electrolyte is a cationic conductor. In addition to the polymer base material, the polymer electrolyte compositions exhibit a conductivity greater than 1×10^{-4} S/cm at 25 °C or below and contain an electrically conductive polymer, a metal salt, a finely divided ionic conductor, and a finely divided inorganic filler material. Certain rechargeable batteries of the invention provide high specific energy (250 to 350 Wh/kg) (gravimetric) and energy density (450 to 550 Wh/l) (volumetric), high cycle life (1000 cycles), low self-discharge and improved safety.

WO 01/17052 A3

INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 00/22917

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 H01M10/40 H01G9/02 H01B1/12

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 H01M H01G H01B

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data, PAJ, CHEM ABS Data, INSPEC, COMPENDEX

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 4 911 995 A (BELANGER ANDRE ET AL) 27 March 1990 (1990-03-27) column 5, line 1 - line 29 column 6, line 22 - line 30 column 8, line 6 - line 15 examples 4,6-8	1,18,19, 24,28, 30,31
A		25,34, 37,38, 66,69, 71,75, 78,79,82
	--- -/--	

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents :

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the international filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the international filing date but later than the priority date claimed

- *T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- *X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- *Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- *&* document member of the same patent family

Date of the actual completion of the international search

21 August 2001

Date of mailing of the international search report

04 09 2001

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040. Tx. 31 651 epo nl,
Fax: (+31-70) 340-3016

Authorized officer

Gamez, A

INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 00/22917

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 5 019 467 A (FUJIWARA YASUO) 28 May 1991 (1991-05-28) column 1, line 66 -column 2, line 2 column 2, line 28 - line 41 column 2, line 62 -column 3, line 13 column 3, line 38 - line 67	28,30,31
A	examples 1,7,8	1,18,24, 25,34, 35,66, 75,82

X	US 5 360 684 A (DUVAL MICHEL ET AL) 1 November 1994 (1994-11-01) column 2, line 54 - line 64 column 5, line 20 - line 23 column 5, line 33 - line 45	28
A	example 3	1,18,24, 25,34,66

X	US 5 100 746 A (GAUTHIER MICHEL ET AL) 31 March 1992 (1992-03-31) column 1, line 64 -column 2, line 10 column 7, line 15 - line 25 column 8, line 3 - line 14	28
A		1,18,24, 25,27, 34,66

X	US 5 415 948 A (GAUTHIER MICHEL ET AL) 16 May 1995 (1995-05-16) column 3, line 3 - line 7 column 4, line 15 - line 27 column 5, line 27 -column 6, line 12 column 10, line 6 - line 43	18,19,24
A	example 3	1,28,30, 31,34, 66, 69-71,75

A	EP 0 741 426 A (YUASA BATTERY CO LTD) 6 November 1996 (1996-11-06) column 3, line 31 - line 35 column 15, line 16 -column 16, line 9 column 16, line 28 - line 57 column 17, line 16 - line 39	1,18,24, 25,27, 28,30, 31,34, 37,38, 66,78

A	US 5 695 873 A (SCANLON JR LAWRENCE G ET AL) 9 December 1997 (1997-12-09) column 1, line 53 -column 2, line 22 column 2, line 50 - line 53	2,4,5, 13,43, 59,65

-/--		

INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 00/22917

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	<p>EP 0 379 372 A (MHB JOINT VENTURE) 25 July 1990 (1990-07-25)</p> <p>page 3, line 2 - line 7 page 3, line 55 -page 4, line 53 page 5, line 41 - line 51</p> <p style="text-align: center;">---</p>	<p>2, 4, 13, 43, 59, 63-65</p>
A	<p>KUMAR B ET AL: "Polymer-ceramic composite electrolytes" JOURNAL OF POWER SOURCES, ELSEVIER SEQUOIA S.A. LAUSANNE, CH, vol. 52, no. 2, 1 December 1994 (1994-12-01), pages 261-268, XP004044634 ISSN: 0378-7753 the whole document</p> <p style="text-align: center;">---</p>	<p>2, 4, 5, 13, 43, 59</p>
A	<p>F. CAPUANO AND AL: "Composite polymer electrolytes" JOURNAL OF ELECTROCHEMICAL SOCIETY, vol. 138, no. 7, 1991, pages 1918-1922, XP002175399 page 1918, right-hand column, paragraph 2 - paragraph 3 page 1920, left-hand column, paragraph 2</p> <p style="text-align: center;">---</p>	<p>2, 4, 5, 13, 43, 59, 64</p>
A	<p>KRAWIEC W ET AL: "Polymer nanocomposites: a new strategy for synthesizing solid electrolytes for rechargeable lithium batteries" JOURNAL OF POWER SOURCES, ELSEVIER SEQUOIA S.A. LAUSANNE, CH, vol. 54, no. 2, 1 April 1995 (1995-04-01), pages 310-315, XP004044428 ISSN: 0378-7753 page 310, left-hand column, paragraph 1 -page 311, left-hand column, paragraph 2 page 312, left-hand column, paragraph 6 -page 312, right-hand column, paragraph 2</p> <p style="text-align: center;">-----</p>	<p>2, 4, 5, 12, 43, 59, 62</p>

INTERNATIONAL SEARCH REPORT

International application No.
PCT/US 00/22917

Box I Observations where certain claims were found unsearchable (Continuation of item 1 of first sheet)

This International Search Report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:
because they relate to subject matter not required to be searched by this Authority, namely:
2. ☐ Claims Nos.:
because they relate to parts of the International Application that do not comply with the prescribed requirements to such an extent that no meaningful International Search can be carried out, specifically:
3. ☐ Claims Nos.:
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box II Observations where unity of invention is lacking (Continuation of item 2 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

see additional sheet

1. ☒ As all required additional search fees were timely paid by the applicant, this International Search Report covers all searchable claims.
2. ☐ As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this International Search Report covers only those claims for which fees were paid, specifically claims Nos.:
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this International Search Report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest.
- ☒ No protest accompanied the payment of additional search fees.

FURTHER INFORMATION CONTINUED FROM PCT/SA/ 210

This International Searching Authority found multiple (groups of) inventions in this international application, as follows:

1. Claims: 1,18-42,66-84

An all-solid laminar electrochemical cell for a battery comprising :

- an anode layer about 0.1-100 μm thick;
- a cathode layer about 0.1-100 μm thick;
- an anode current collector about 0.5-50 μm thick attached to said anode;
- a cathode current collector about 0.5-50 μm thick attached to said cathode; and a layer of solid polymer electrolyte disposed between said anode and cathode layers.

2. Claims: 2-17, 43-65

An all-solid state laminar electrochemical cell for a battery comprising :

- an anode layer;
- a cathode layer ;
- an anode current collector attached to said anode;
- a cathode current collector attached to said cathode; and
- a layer of solid polymer electrolyte disposed between said anode and cathode layers, said polymer electrolyte being a cationic conductor having a conductivity of at least 10^{-4} S/cm at 25 C and comprising a mixture of
 - a base polymer material comprising at least one ionically conductive polymer and having an initial conductivity of at least about 1×10^{-5} S/cm at 25 C when combined with a metal salt;
 - a metal salt;
 - an inorganic filler having an average particle size $< 0.05 \mu\text{m}$ in diameter and a surface area of at least about $100 \text{ m}^2/\text{g}$; and
 - an ionic conducting material having an average particle size $< 0.1 \mu\text{m}$ in diameter and an initial conductivity of at least 2×10^{-3} S/cm at 25 C.

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/US 00/22917

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
US 4911995	A	27-03-1990	CA 1288473 A	03-09-1991
			AT 76702 T	15-06-1992
			DE 3871417 A	02-07-1992
			EP 0285476 A	05-10-1988
			ES 2032983 T	01-03-1993
			GR 3005547 T	07-06-1993
			JP 63241859 A	07-10-1988
			US 4824746 A	25-04-1989
US 5019467	A	28-05-1991	JP 1130473 A	23-05-1989
			JP 2692816 B	17-12-1997
US 5360684	A	01-11-1994	CA 2109247 A	27-04-1995
			EP 0651453 A	03-05-1995
			AT 152293 T	15-05-1997
			DE 69310177 D	28-05-1997
			DE 69310177 T	31-07-1997
US 5100746	A	31-03-1992	FR 2616970 A	23-12-1988
			AT 89102 T	15-05-1993
			CA 1340407 A	23-02-1999
			DE 3880747 A	09-06-1993
			DE 3880747 T	11-11-1993
			EP 0298800 A	11-01-1989
			ES 2041817 T	01-12-1993
			WO 8810519 A	29-12-1988
			JP 1503660 T	07-12-1989
			JP 3047355 B	29-05-2000
US 5415948	A	16-05-1995	CA 2051614 A	18-03-1993
			AT 152858 T	15-05-1997
			DE 69219542 D	12-06-1997
			DE 69219542 T	06-11-1997
			EP 0533577 A	24-03-1993
			JP 5314986 A	26-11-1993
EP 0741426	A	06-11-1996	JP 3236157 A	22-10-1991
			JP 4043556 A	13-02-1992
			JP 2861329 B	24-02-1999
			JP 4104461 A	06-04-1992
			CA 2036098 A	14-08-1991
			DE 69127109 D	11-09-1997
			DE 69127109 T	22-01-1998
			DE 69132176 D	08-06-2000
			DE 69132176 T	09-11-2000
			EP 0442681 A	21-08-1991
			US 5290592 A	01-03-1994
US 5695873	A	09-12-1997	NONE	
EP 0379372	A	25-07-1990	US 4990413 A	05-02-1991
			DE 69003692 D	11-11-1993
			DE 69003692 T	11-05-1994
			ES 2047838 T	01-03-1994
			HK 1008402 A	07-05-1999
			JP 2236906 A	19-09-1990
			JP 3055144 B	26-06-2000

